

Chemical Basis of Life

Vocabulary

Polarity

Concentration

4 carbon compounds

monosaccharide

lipid

enzyme

Hydrogen bond

Solution

monomer

polysaccharide

fatty acid

catalyst

Adhesion

Solute

polymer

amino acid

nucleic acid

pH

Cohesion

Solvent

carbohydrate

protein

nucleotide

specific heat

Concepts to Know

WATER

Why so special and important?

Nearly 75% of our earth is covered in water.

Our bodies are 60-65% water. Our blood is 85% water.

None of the many reactions that occur in our bodies would occur without water.

Composition of water: H_2O

Properties of Water

- Water is **POLAR**.

The diagram shows that one side of the water molecule is **positively charged** and the other is **negatively charged**.

The molecule itself is neutral. These opposite charges make water a **polar molecule**.

- Water undergoes **ADHESION** and **COHESION** due to its polarity.

The **negative oxygen** of one water molecule is attracted to the **positive hydrogen** of another molecule forming a **hydrogen bond**. In other words water likes to stick to itself. Water sticking to water is called **cohesion**. Water's polarity allows it to attract to other surfaces too (like glass). Water sticking to something else is called **adhesion**.

1. In the diagram to the right use dotted lines to draw in the bonds that form between water molecules.

Positives connect to negatives, up to four bonds per water molecule.

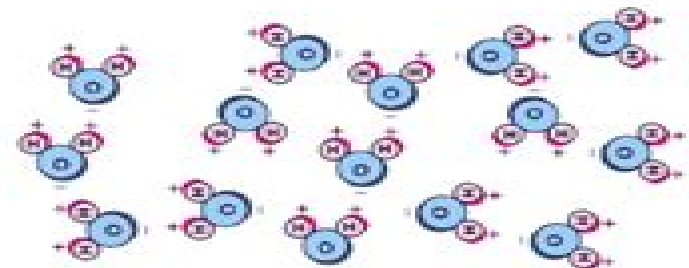
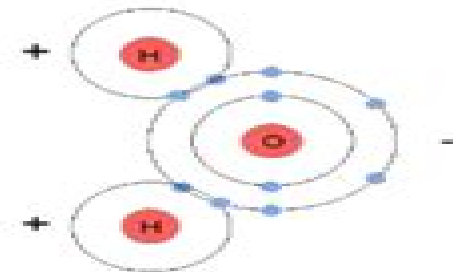
2. What is the name of this type of bond?

Hydrogen or "H" bond

3. What special property do the bonds give to water.

Several: cohesion first and foremost (water sticking to itself).

Creates surface tension on the water (water skipper can seemingly "walk" on water)



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