

pH and Buffers

ACIDS AND BASES

Bronsted and Lowry have defined acids as proton donors and bases as proton acceptors. Each acid, therefore, has always a conjugated base,



Alkali refers for the compound that yields hydroxyl ions on dissociation.



The strength of acids or bases is the ability to lose H^+ or OH^- . Weak acids and bases dissociate only to limited extent.



Potassium hydroxide



(Weak acid)



(Weak base) (Aniline)

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