

Chemical Bonds Study Guide

Introduction

Chemical bonds are the forces that hold atoms together in a molecule or compound. Understanding chemical bonding is essential to understanding the behavior of matter in various fields of science. In this study guide, we will cover the different types of chemical bonds and their characteristics.

Types of Chemical Bonds

Covalent Bonds

Covalent bonds occur when atoms share one or more pairs of electrons. The sharing of electrons allows each atom to fill its outermost energy level and become more stable. Covalent bonds can be polar or nonpolar, depending on the electronegativity difference between the atoms.

Ionic Bonds

Ionic bonds occur when one or more electrons are transferred from one atom to another, resulting in the formation of charged ions. The attraction between the opposite charges holds the ions together in a crystal lattice structure.

Metallic Bonds

Metallic bonds occur when metal atoms release their valence electrons to a common pool of electrons. This shared electron sea holds the metal cations together in a crystal lattice structure, which can move freely through the lattice, allowing for electrical conductivity and malleability.

Bonding Theories

Valence Shell Electron Pair

Valence shell electron pair theory (VSEPR) predicts the shape of a molecule based on the arrangement of valence electrons in a molecule or compound. This is done by considering the repulsion between electron pairs, which can be bonding or non-bonding (lone pairs).

Molecular Orbital

Molecular orbital theory (MOT) describes the chemical bonding between atoms by considering the overlap of atomic orbitals to form molecular orbitals. It provides the energy of a molecule based on the arrangement of its electron pairs.

Hybridization

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John A. Olmsted, Gregory M. Williams



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