

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Match the symbol on the left with its appropriate description on the right.

<u> d </u>	(a) A precipitate forms.
<u> a </u> ↓	(b) A gas forms.
<u> b </u>	(c) A reversible reaction occurs.
<u> f </u> (l)	(d) Heat is applied to the reactants.
<u> e </u> (aq)	(e) A chemical is dissolved in water.
<u> c </u>	(f) A chemical is in the liquid state.

2. Finish balancing the equation:



3. In each of the following formulas, write the total number of atoms present.

12 atoms a. 4SO_2

16 atoms b. 3O_2

51 atoms c. $3\text{Al}_2(\text{SO}_4)_3$

3×10^{24} atoms d. $6 \times 10^3 \text{HNO}_3$

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride → aluminum fluoride + copper metal



5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



6. a. Balance the following equation:



Chemical Equations And Reactions Section 1 Answers

**John Obimakinde, Samuel
Obimakinde, Ebenezer
Obimakinde, Fredrick Akinbolade**

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