

## Equilibrium Constant - Worksheet 1\_ANSWERS

1. Write the equilibrium expression for the following reactions:



$$K = \frac{[U]^u [V]^v}{[W]^w [X]^x}$$



$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$



$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$



$$K = \frac{[HCl]^2}{[H_2][Cl_2]}$$



$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$



$$K = \frac{[Zn^{2+}]}{[Ag^+]^2}$$

\*Note: solids are not included



$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:



The  $K_{eq}$  for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50 mol/L. Assuming  $H_2$  and  $Br_2$  are present in equal amounts, calculate the concentration of  $H_2$  at equilibrium.

$$[H_2] = 0.50 \text{ mol/L}$$

# Chemical Equilibrium Answers

**James E. Banks**



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