

Question 7.4:

Write the expression for the equilibrium constant, K_c for each of the following reactions:



Answer:

(i)
$$K_c = \frac{[\text{NO}_{\text{(g)}}]^2 [\text{Cl}_{2\text{(g)}}]}{[\text{NOCl}_{\text{(g)}}]^2}$$

(ii)
$$K_c = \frac{[\text{CuO}_{\text{(s)}}]^2 [\text{NO}_{2\text{(g)}}]^4 [\text{O}_{2\text{(g)}}]}{[\text{Cu(NO}_3)_2\text{(s)}]^2}$$

$$= [\text{NO}_{2\text{(g)}}]^4 [\text{O}_{2\text{(g)}}]$$

(iii)
$$K_c = \frac{[\text{CH}_3\text{COOH}_{\text{(aq)}}][\text{C}_2\text{H}_5\text{OH}_{\text{(aq)}}]}{[\text{CH}_3\text{COOC}_2\text{H}_5\text{(aq)}][\text{H}_2\text{O}_{\text{(l)}}]} = \frac{[\text{CH}_3\text{COOH}_{\text{(aq)}}][\text{C}_2\text{H}_5\text{OH}_{\text{(aq)}}]}{[\text{CH}_3\text{COOC}_2\text{H}_5\text{(aq)}]}$$

(iv)
$$K_c = \frac{[\text{Fe(OH)}_{3\text{(s)}}]}{[\text{Fe}^{3+}_{\text{(aq)}}][\text{OH}^-_{\text{(aq)}}]^3}$$

$$= \frac{1}{[\text{Fe}^{3+}_{\text{(aq)}}][\text{OH}^-_{\text{(aq)}}]^3}$$

(v)
$$K_c = \frac{[\text{IF}_5]^2}{[\text{I}_{2\text{(s)}}][\text{F}_2]^5}$$

$$= \frac{[\text{IF}_5]^2}{[\text{F}_2]^5}$$

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