

Equilibrium Constant - Worksheet 1_ANSWERS

1. Write the equilibrium expression for the following reactions:



$$K = \frac{[U]^u [V]^v}{[W]^w [X]^x}$$



$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$



$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$



$$K = \frac{[HCl]^2}{[H_2][Cl_2]}$$



$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$



$$K = \frac{[Zn^{2+}]}{[Ag^+]^2}$$

*Note: solids are not included



$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:



The K_{eq} for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50 mol/L. Assuming H_2 and Br_2 are present in equal amounts, calculate the concentration of H_2 at equilibrium.

$$[H_2] = 0.50 \text{ mol/L}$$

Chemical Equilibrium Answers Key

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1. Understanding the eBook Chemical Equilibrium Answers Key
 - The Rise of Digital Reading Chemical Equilibrium Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equilibrium Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Equilibrium Answers Key
 - Personalized Recommendations
 - Chemical Equilibrium Answers Key User Reviews and Ratings
 - Chemical Equilibrium Answers Key and Bestseller Lists
5. Accessing Chemical Equilibrium Answers Key Free and Paid eBooks

- Chemical Equilibrium Answers Key Public Domain eBooks
 - Chemical Equilibrium Answers Key eBook Subscription Services
 - Chemical Equilibrium Answers Key Budget-Friendly Options
6. Navigating Chemical Equilibrium Answers Key eBook Formats
- ePub, PDF, MOBI, and More
 - Chemical Equilibrium Answers Key Compatibility with Devices
 - Chemical Equilibrium Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemical Equilibrium Answers Key
 - Highlighting and Note-Taking Chemical Equilibrium Answers Key
 - Interactive Elements Chemical Equilibrium Answers Key
8. Staying Engaged with Chemical Equilibrium Answers Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Answers Key
9. Balancing eBooks and Physical Books Chemical Equilibrium Answers Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Answers Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equilibrium Answers Key
- Setting Reading Goals Chemical Equilibrium Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Equilibrium Answers Key
- Fact-Checking eBook Content of Chemical Equilibrium Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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