

16 • Chemical Equilibrium

1. Consider the reaction system,



The equilibrium constant expression is

- a) $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$ d) $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$
 b) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$ e) $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$
 c) $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{H}_2]}$

2. Given the equilibrium,

if this equilibrium is established by beginning with equal number of moles of SO_2 and O_2 in a 1.0 Liter bulb, thenthe following *must* be true at equilibrium:

- a) $[\text{SO}_2] = [\text{SO}_3]$ d) $[\text{SO}_2] < [\text{O}_2]$
 b) $2[\text{SO}_2] = 2[\text{SO}_3]$ e) $[\text{SO}_2] > [\text{O}_2]$
 c) $[\text{SO}_2] = [\text{O}_2]$

Questions 3 & 4 refer to the following:

At a given temperature, 0.300 mole NO , 0.200 mol Cl_2 and 0.500 mol ClNO were placed in a 25.0 Liter container. The following equilibrium is established: $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{(g)}$

3. At equilibrium, 0.600 mol of
- ClNO
- was present. The number of
- moles*
- of
- Cl_2
- present at equilibrium is

- a) 0.050 d) 0.200
 b) 0.100 e) 0.250
 c) 0.150

P R A C T I C E T E S T

4. The equilibrium constant,
- K_c
- , is:

- a) 4.45×10^{-4} d) 0.167
 b) 6.67×10^{-4} e) 1500
 c) 0.111

5. At
- 985°C
- , the equilibrium constant for the reaction,



is 1.63. What is the equilibrium constant for the reverse reaction?

- a) 1.63 d) 0.613
 b) 0.815 e) 1.00
 c) 2.66

6. What is the relationship between
- K_p
- and
- K_c
- for the reaction,
- $2\text{HCl(g)} \rightleftharpoons \text{H}_2\text{(g)} + \text{Cl}_2\text{(g)}$
- ?

- a) $K_p = K_c(RT)^{-1}$ d) $K_p = K_c$
 b) $K_p = K_c(RT)$ e) $K_p = K_c(2RT)$
 c) $K_p = K_c(RT)^2$

7. For the reaction
- $2\text{NO}_2\text{(g)} \rightleftharpoons \text{N}_2\text{O}_4\text{(g)}$
- ,
- K_p
- at
- 25°C
- is 7.3, when all partial pressures are expressed in atmospheres. What is
- K_c
- for this reaction? [
- $R=0.0821 \text{ L}\cdot\text{atm}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$
-]

- a) 4270 d) 179
 b) 0.0119 e) 2.06
 c) 0.291

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D Keegan



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Table of Contents Chemical Equilibrium Practice Test Answers South Pasadena

1. Understanding the eBook Chemical Equilibrium Practice Test Answers South Pasadena
 - The Rise of Digital Reading Chemical Equilibrium Practice Test Answers South Pasadena
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Equilibrium Practice Test Answers South Pasadena
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Equilibrium Practice Test Answers South Pasadena
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemical Equilibrium Practice Test Answers South Pasadena
 - Personalized Recommendations
 - Chemical Equilibrium Practice Test Answers South Pasadena User Reviews and Ratings
 - Chemical Equilibrium Practice Test Answers South Pasadena and Bestseller Lists
5. Accessing Chemical Equilibrium Practice Test Answers South Pasadena Free and Paid eBooks
 - Chemical Equilibrium Practice Test Answers South Pasadena Public Domain eBooks
 - Chemical Equilibrium Practice Test Answers South Pasadena eBook Subscription Services
 - Chemical Equilibrium Practice Test Answers South Pasadena Budget-Friendly Options
6. Navigating Chemical Equilibrium Practice Test Answers South Pasadena eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Equilibrium Practice Test Answers South Pasadena Compatibility with Devices
 - Chemical Equilibrium Practice Test Answers South Pasadena Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Equilibrium Practice Test Answers South Pasadena
 - Highlighting and Note-Taking Chemical Equilibrium Practice Test Answers South Pasadena
 - Interactive Elements Chemical Equilibrium Practice Test Answers South Pasadena
8. Staying Engaged with Chemical Equilibrium Practice Test Answers South Pasadena
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Equilibrium Practice Test Answers South Pasadena
9. Balancing eBooks and Physical Books Chemical Equilibrium Practice Test Answers South Pasadena
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Equilibrium Practice Test Answers South Pasadena
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Equilibrium Practice Test Answers South Pasadena
 - Setting Reading Goals Chemical Equilibrium Practice Test Answers South Pasadena
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemical Equilibrium Practice Test Answers South Pasadena
 - Fact-Checking eBook Content of Chemical Equilibrium Practice Test Answers South Pasadena
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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