

Chemical equilibrium

2. The concentration of a pure solid or liquid phase is not included in the expression of equilibrium constant because
- (a) solid and liquid concentrations are independent of their quantities.
 - (b) solid and liquids react slowly.
 - (c) solid and liquids at equilibrium do not interact with gaseous phase.
 - (d) the molecules of solids and liquids cannot migrate to the gaseous phase.
3. For the reversible reaction, $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ at 500°C , the value of K_p is 1.44×10^{-5} when partial pressure is measured in atmospheres. The corresponding value of K_c , with concentration in mole litre⁻¹ is
- | | |
|---|---|
| (a) $\frac{1.44 \times 10^{-5}}{(0.082 \times 500)^{-2}}$ | (b) $\frac{1.44 \times 10^{-5}}{(8.314 \times 773)^{-2}}$ |
| (c) $\frac{1.44 \times 10^{-5}}{(0.082 \times 773)^2}$ | (d) $\frac{1.44 \times 10^{-5}}{(0.082 \times 773)^{-2}}$ |
4. When two reactants, A and B are mixed to give products C and D, the reaction quotient Q at the initial stages of the reaction
- (a) is zero.
 - (b) decreases with time.
 - (c) is independent of time.
 - (d) increases with time.

Chemical Equilibrium Worked Exam Question

S Ben Porath



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