# Chapter Test B

## Chapter: Chemical Formulas and Chemical Compounds

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Changing a subscript in a correctly written chemical formula will
  - a. change the electron configuration of that element.
  - b. change the charges on the other ions in the compound.
  - c. change the formula so that it no longer represents the same compound.
  - d. have no effect on the formula.
- 2. Using the Stock system of nomenclature, Cr<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> is named
  - a. chromium(II) sulfate.
  - b. chromic sulfate.
  - c. dichromium trisulfate.
  - chromium(III) sulfate.
- \_\_\_\_\_\_ 3. In a polyatomic ion, the algebraic sum of the oxidation numbers of all atoms is equal to
  - a. 0.
  - ь. 10.
  - c. the number of atoms in the ion.
  - d. the charge on the ion.
- \_\_\_\_\_ 4. The first part of the name of a binary ionic compound is the
  - cation.
  - b. polyatomic ion.
  - c. oxyanion.
  - d. anion.
- \_\_\_\_\_5. The empirical formula may not represent the actual composition of a(n)
  - ionic compound.
  - b. crystal.
  - c. sailt.
  - d. molecular compound.
- \_\_\_\_\_ 6. The formula for phosphoric acid is
  - a. HPO<sub>4</sub>.
  - b. H(PO<sub>4</sub>)<sub>2</sub>.
  - c. HaPOn
  - d. H<sub>3</sub>PO<sub>4</sub>.

# **Chemical Formulas And Chemical Compounds Assessment**

Roy M. Mortier, Malcolm F. Fox, Stefan Orszulik

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