

## 9

## CHEMICAL NAMES AND FORMULAS

**SECTION 9.1 NAMING IONS (pages 253–258)**

*This section explains the use of the periodic table to determine the charge of an ion. It also defines polyatomic ion and gives the names and formulas for the most common polyatomic ions.*

**► Monatomic Ions (pages 253–256)**

1. What are monatomic ions?

Monatomic ions are ions consisting of only one atom.

2. How is the ionic charge of a Group 1A, 2A, or 3A ion determined?

The ionic charge is numerically equal to the group number.

3. How is the ionic charge of a Group 5A, 6A, or 7A ion determined?

The charge of an ion in Groups 5A, 6A, or 7A is determined by subtracting 8 from the group number.

4. Circle the letter of the type of element that often has more than one common ionic charge.

- a. alkali metal
- b. alkaline earth metal
- ☒ c. transition metal
- d. nonmetal

5. The Stock system of naming transition metal cations uses a Roman numeral in parentheses to indicate the numeric value of the ionic charge.

6. An older naming system uses the suffix *-ous* to name the cation with the lesser charge, and the suffix *-ic* to name the cation with the greater charge.

7. What is a major advantage of the Stock system over the old naming system?

The Stock system gives the actual charge of the ion.

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## **Table of Contents Chemical Names And Formulas Chapter 9 Quiz**

1. Understanding the eBook Chemical Names And Formulas Chapter 9 Quiz
  - The Rise of Digital Reading Chemical Names And Formulas Chapter 9 Quiz
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Names And Formulas Chapter 9 Quiz
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Names And Formulas Chapter 9 Quiz
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Names And Formulas Chapter 9 Quiz
  - Personalized Recommendations
  - Chemical Names And Formulas Chapter 9 Quiz User Reviews and Ratings
  - Chemical Names And Formulas Chapter 9 Quiz and Bestseller Lists

5. Accessing Chemical Names And Formulas Chapter 9 Quiz Free and Paid eBooks
  - Chemical Names And Formulas Chapter 9 Quiz Public Domain eBooks
  - Chemical Names And Formulas Chapter 9 Quiz eBook Subscription Services
  - Chemical Names And Formulas Chapter 9 Quiz Budget-Friendly Options
6. Navigating Chemical Names And Formulas Chapter 9 Quiz eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Names And Formulas Chapter 9 Quiz Compatibility with Devices
  - Chemical Names And Formulas Chapter 9 Quiz Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Names And Formulas Chapter 9 Quiz
  - Highlighting and Note-Taking Chemical Names And Formulas Chapter 9 Quiz
  - Interactive Elements Chemical Names And Formulas Chapter 9 Quiz
8. Staying Engaged with Chemical Names And Formulas Chapter 9 Quiz
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Names And Formulas Chapter 9 Quiz
9. Balancing eBooks and Physical Books Chemical Names And Formulas Chapter 9 Quiz
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Names And Formulas Chapter 9 Quiz
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Names And Formulas Chapter 9 Quiz
  - Setting Reading Goals Chemical Names And Formulas Chapter 9 Quiz
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Names And Formulas Chapter 9 Quiz
  - Fact-Checking eBook Content of Chemical Names And Formulas Chapter 9 Quiz
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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