

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary term.

- | | |
|--|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²³ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of a compound is a compound _____ molar mass | |
| _____ the number representing 6.02 x 10 ²³ representative particles of a substance _____ Avogadro's number | |
| _____ 1 mole = 6.02 ²³ particles _____ empirical formula | |
| _____ the number representing the number of particles of a substance is a substance _____ 22.4 L _____ Avogadro's number _____ standard temperature (STP, 0°C) | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following section by displaying knowledge of the terms and your understanding of the concepts involved with the terms of the definitions corresponding to each (don't give) or explain.

Measuring whole units of something, or mass, and its volume for a single quantity called the _____ mole. _____ Representative representative particles for 1 mole of a substance is _____ Avogadro's number _____

It is the mass of particles of a compound, substance and together the _____ percent composition _____ of the atoms making the compound. Other two techniques for measuring the mass of a substance are _____ molar mass _____ of the compound. These are _____ 6.02 x 10²³ _____ representative particles the mass of one substance.

True or False

_____ 1 mole of any substance has a mass of 1 gram.

_____ 1 mole of substance contains 6.02 x 10²³ atoms.

_____ 1 mole of representative particles of a substance is the substance. **True**

_____ 1 mole of 10²³ is composed of one mole of carbon and four atoms of oxygen.

_____ 1 mole mass of nitrogen gas (N₂) is **28.02 g**.

Problems

Chemical Quantities Vocabulary Review Answer Key

PT Brinkman



Chemical Quantities Vocabulary Review Answer Key:

Decoding **Chemical Quantities Vocabulary Review Answer Key**: Revealing the Captivating Potential of Verbal Expression

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