#### Types of Chemical Reactions Worksheet #1

#### HI/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 2Mg + O<sub>2</sub> → 2 MgO Synthesis
- Cal<sub>2</sub> + Cl<sub>2</sub> → CaCl<sub>2</sub> + I<sub>2</sub> Single Replacement
- 3) 3KOH + AlCl<sub>3</sub> → Al(OH)<sub>3</sub> + 3KCl Double Replacement
- C + O<sub>2</sub> → CO<sub>2</sub> Synthesis
- \*5) Ca(OH)<sub>2</sub> + 2HC1 → CaCl<sub>2</sub> + 2H<sub>2</sub>O Double Replacement
- 6) 2KClO<sub>3</sub> → 2KCl+3 O<sub>2</sub> Decomposition
- 7) 3CuSO<sub>4</sub> + 2Al → Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + 3Cu Single Replacement
- 8) Na<sub>2</sub>S + 2AgNO<sub>1</sub> → 2NaNO<sub>1</sub> + Ag<sub>2</sub>S Double Replacement
- 9)  $3H_2 + N_2 \rightarrow 2NH_3$  Synthesis
- \*10) 2Na + 2H2O -> 2NaOH + H2 Single Replacement

\*Remember: "HyO" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na<sub>2</sub>CrO<sub>4</sub> + PbCl<sub>2</sub> → 2 NaCl + PbCrO<sub>4</sub>
- Single replacement: Cl<sub>2</sub> + 2 NaBr → 2 NaCl + Br<sub>2</sub>
- Decomposition: Mg(ClO<sub>3</sub>)<sub>2</sub> → MgCl<sub>2</sub> + 3 O<sub>2</sub>
- 14) Synthesis: 2 H<sub>2</sub> + O<sub>2</sub> → 2 H<sub>2</sub>O
- 15) Double replacement: 3 Ca(OH)<sub>2</sub> + 2 FeCl<sub>3</sub> → 2 Fe(OH)<sub>1</sub> + 3 CaCl<sub>2</sub>
- 16) Single replacement: Fe + Cu(NO<sub>3</sub>)<sub>2</sub> → Fo(NO<sub>1</sub>)<sub>2</sub> + Cu [Assume Fe<sup>2+</sup>]
- Decomposition: 2 Hg<sub>2</sub>O → 4 Hg + O<sub>2</sub>
- 18) Synthesis:  $S + O_2 \rightarrow SO_2$
- Double replacement: AgNO<sub>3</sub> + KI → AgI + KNO<sub>3</sub>
- 20) Single replacement: Cu + 2 AgNO<sub>3</sub> → 2 Ag + Cu(NO<sub>3</sub>)<sub>2</sub> [Copper (II) is used here]

# **Chemical Reactions Answers**

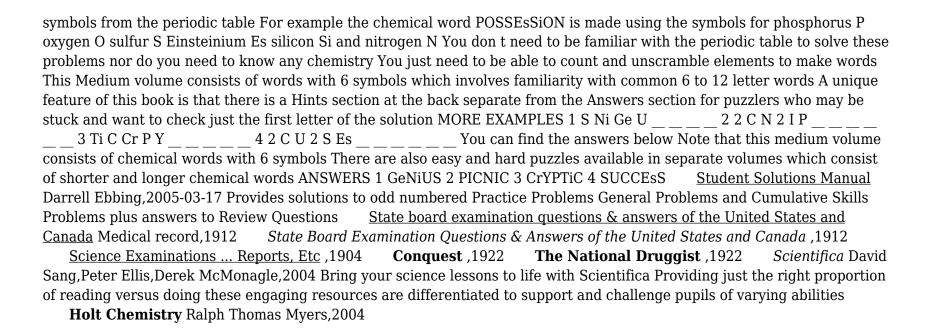
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#### **Table of Contents Chemical Reactions Answers**

- 1. Understanding the eBook Chemical Reactions Answers
  - The Rise of Digital Reading Chemical Reactions Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemical Reactions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Reactions Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemical Reactions Answers
  - Personalized Recommendations
  - Chemical Reactions Answers User Reviews and Ratings
  - Chemical Reactions Answers and Bestseller Lists

- 5. Accessing Chemical Reactions Answers Free and Paid eBooks
  - Chemical Reactions Answers Public Domain eBooks
  - Chemical Reactions Answers eBook Subscription Services
  - Chemical Reactions Answers Budget-Friendly Options
- 6. Navigating Chemical Reactions Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Chemical Reactions Answers Compatibility with Devices
  - Chemical Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Reactions Answers
  - Highlighting and Note-Taking Chemical Reactions Answers
  - Interactive Elements Chemical Reactions Answers
- 8. Staying Engaged with Chemical Reactions Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Reactions Answers
- 9. Balancing eBooks and Physical Books Chemical Reactions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Reactions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Reactions Answers
  - Setting Reading Goals Chemical Reactions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Reactions Answers
  - Fact-Checking eBook Content of Chemical Reactions Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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