



1. A compound X contains carbon, hydrogen and oxygen only.

X contains 54.54% of carbon by mass, 9.09% of hydrogen by mass and 36.37% of oxygen by mass.

- (i) Calculate the empirical formula of compound X.



[2]

- (ii) Compound X has a relative molecular mass of 88.

Deduce the molecular formula of compound X.



[2]

Iodine forms an oxide which has the composition by mass: I, 76.0%; O, 24.0%.

- (iii) Use this information to determine the empirical formula of this oxide of iodine.

empirical formula ..... [3]



# Chemical Stoichiometry Test

**Thomas Evangelist**



## **Chemical Stoichiometry Test:**

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