#### CHE 102 Final Exam Study Guide

#### Part One: The Mole & Stoichiometry

1. Balance the following equations:

a.  $\__HG + \__Zn \rightarrow \__H_2 + \__ZnG_2$ 

b. \_\_\_N<sub>2</sub>H<sub>4</sub> + \_\_\_H<sub>2</sub>O<sub>2</sub>  $\Rightarrow$  \_\_\_N<sub>2</sub> + \_\_\_H<sub>2</sub>O

c.  $__C_7H_{16} + __O_2 \rightarrow __CO_2 + __H_2O$ 

d.  $\underline{\hspace{1cm}}$  Ca(NO<sub>3</sub>)<sub>2</sub>  $\rightarrow$   $\underline{\hspace{1cm}}$  CaCO<sub>3</sub> +  $\underline{\hspace{1cm}}$  NaNO<sub>3</sub>

2. How does the Law of Conservation of Mass apply to balancing chemical equations?

 A recipe calls for 4.0 grams (approximately 1 teaspoon) of cream of tartar. How many moles of cream of tartar (KC<sub>a</sub>H<sub>3</sub>O<sub>b</sub>) is this? Show your work!

 An apple contains 12.6 grams of fructose, C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>. How many molecules of fructose is this? Show your work!

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**JS Bruner** 

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