

MIDTERM PREPARATION PROBLEMS LABORATORY

Winter 2019

This problems laboratory is designed to help you prepare for the midterm. Those who have been exempted from the laboratory this semester may attend ANY of the problem laboratories running in week 5.

1. A 0.5693 g sample of solid lactic acid, $\text{HC}_3\text{H}_5\text{O}_3$ (MM = 90.08 g/mol) is burned in a bomb calorimeter in the presence of excess oxygen to produce carbon dioxide gas and water vapour. The bomb calorimeter was surrounded by 1.00 kg water, and the temperature rose from 23.10 °C to 24.95 °C. The heat capacity of the entire calorimeter (bomb and water bath) was given as 4.985 kJ K⁻¹. Determine the enthalpy change for 1.00 mole of lactic acid at 25 °C based on these conditions.
2. In a calorimetry experiment, 50.0 g of a 2.04 mol/kg HCl solution is mixed with 50.0 g of a 2.13 mol / kg NaOH solution. The specific heat of the resulting solution is 3.90 J °C⁻¹g⁻¹. The temperature of the mixture rises from 22.1 °C to 37.0 °C. Calculate the enthalpy change for the overall reaction.
3. Our bodies can use glucose, $\text{C}_6\text{H}_{12}\text{O}_6$ (s) (MM = 180.2 g/mol), as a fuel. The heat of combustion of glucose is 2.820×10^3 kJ/mol. The human body has a specific heat of 4.18 J °C⁻¹g⁻¹. How many grams of glucose would an 80.0 kg person need to consume to raise the body temperature from 36.0 to 39.0 °C? Assume that all energy goes to warm the body.
4. A cylinder containing 12.0 L of an ideal gas at a pressure of 25.0 atm expands against an external pressure of 1.00 atm until the final pressure of the gas is 1.00 atm. Calculate the work done by the system.
5. Write the thermochemical equation for each of the following:
 - a) Heat of formation of ammonium iodide
 - b) $\Delta H = -2$ [enthalpy of formation of mercury (I) bromate, $\text{Hg}_2(\text{BrO}_3)_2$]
 - c) Heat of combustion of benzoic acid, $\text{C}_7\text{H}_6\text{O}_2$ (s)
6. Given the reaction enthalpies:

$$2 \text{P(s)} + 3 \text{Cl}_2(\text{g}) \rightarrow 2 \text{PCl}_3(\text{g}) ; \Delta H = -574 \text{ kJ}$$

$$2 \text{P(s)} + 5 \text{Cl}_2(\text{g}) \rightarrow 2 \text{PCl}_5(\text{l}) ; \Delta H = -887 \text{ kJ}$$
 Calculate the reaction enthalpy for the reaction: $\text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g}) \rightarrow \text{PCl}_5(\text{l})$
7. Calculate the standard reaction enthalpy for the reaction $3 \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_6\text{H}_6(\text{l})$ given the standard enthalpies of combustion of $\text{C}_2\text{H}_2(\text{g})$ and $\text{C}_6\text{H}_6(\text{l})$ are -1300.0 kJ / mole and -3268.0 kJ / mol , respectively.
8. Given :
 - Br – Br bond energy = 193 kJ / mol ;
 - Cl – Cl bond energy = 242 kJ / mol ;
 - Br – Cl bond energy = 218 kJ / mol ;
 Heat of vaporization of bromine = 31 kJ / mol ;

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**Frederick A. Bettelheim, Joseph M.
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