

- A) They are restricted to metals. C) Two reactants produce two products.
 B) They involve a single product. D) Any metal replaces any other metal.

16. State the type of reaction: $2\text{H}_2 + 2\text{H}_2\text{O} \rightarrow 2\text{H}_2\text{O}_2 + \text{H}_2$

- A) synthesis B) decomposition C) single displacement
 D) double displacement E) combustion F) no reaction

17. State the type of reaction: $4\text{Cr} + 3\text{O}_2 \rightarrow 2\text{Cr}_2\text{O}_3$

- A) synthesis B) decomposition C) single displacement
 D) double displacement E) no reaction

18. Determine the number of moles present in 30.5-g aluminum chloride.

- A) 0.344 mol B) 4.10 mol C) 1.21 mol D) 700 mol

19. What is the mass in grams of 1.00×10^{24} atoms of oxygen as O_2 ?

- A) 0.132×10^3 g B) 0.269×10^3 g C) 0.30×10^3 g D) 0.30×10^3 g

20. The number of molecules in 2.0 mol of carbon dioxide is:

- A) 1.0×10^{24} B) 3.0×10^{24} C) 6.0×10^{24} D) 3.01×10^{24} E) 1.2×10^{24}

21. What is the percentage composition of COF?

- A) 50% C, 50% O B) 12% C, 88% O C) 25% C, 75% O D) 40% C, 60% O

22. The mole ratio of carbon to hydrogen to oxygen in a compound is 1 mol C : 2 mol H : 1 mol O. What is the empirical formula of the compound?

- A) CHO B) CH₂O C) C₂H₄O₂ D) C₂H₄O₄

23. The molecular formula of a compound whose empirical formula is CH₂O₂ and molar mass is 284 g/mol is:

- A) CH₂O₂ B) C₂H₄O₄ C) C₃H₆O₆ D) C₄H₈O₈ E) C₅H₁₀O₁₀

24. Lithium and nitrogen react in a combination reaction to produce lithium nitride:



In a particular experiment, 3.50-g samples of each reactant are reacted. The theoretical yield of lithium nitride is _____ g.

- A) 3.52 B) 2.89 C) 17.6 D) 5.85 E) 8.7

25. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of KClO₃ required to produce 5.0 mol of oxygen is:

- A) 1.42 B) 0.422 C) 2.08 D) 4.38 E) 3.33

26. The theoretical yield of copper was calculated to be 288.0 g and the % yield was 80.90%. Determine the actual yield.

- A) 28.47 g B) 264.3 g C) 284.7 g D) 205.1 g

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Darrell D. Ebbing, David Bookin



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