

Chemistry

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(Chapter - 3) (Chemical Kinetics) (Exercises)
(Class 12)

Therefore,

$$\text{Rate} = \frac{1}{2} \frac{d[\text{NH}_3]}{dt} = \frac{d[\text{N}_2]}{dt} = \frac{1}{3} \frac{d[\text{H}_2]}{dt}$$

However, it is given that the reaction is of zero order reaction. Therefore,

$$k = \frac{1}{2} \frac{d[\text{NH}_3]}{dt} = \frac{d[\text{N}_2]}{dt} = \frac{1}{3} \frac{d[\text{H}_2]}{dt} = 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ s}^{-1}$$

Therefore, the rate of production of N_2 is

$$\frac{d[\text{N}_2]}{dt} = 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ s}^{-1}$$

And, the rate of production of H_2 is

$$\frac{d[\text{H}_2]}{dt} = 3 \times 2.5 \times 10^{-4} \text{ mol L}^{-1} \text{ s}^{-1} = 7.5 \times 10^{-4} \text{ mol L}^{-1} \text{ s}^{-1}$$

Question 3.4:

The decomposition of dimethyl ether leads to the formation of CH_4 , H_2 and CO and the reaction rate is given by, $\text{Rate} = k[\text{CH}_3\text{OCH}_3]^{3/2}$

The rate of reaction is followed by increase in pressure in a closed vessel, so the rate can also be expressed in terms of the partial pressure of dimethyl ether, i.e., $\text{Rate} = k(P_{\text{CH}_3\text{OCH}_3})^{3/2}$

If the pressure is measured in bar and time in minutes, then what are the units of rate and rate constants?

Answer 3.4:

If pressure is measured in bar and time in minutes, then

Unit of rate = bar min^{-1}

$$\text{Rate} = k(P_{\text{CH}_3\text{OCH}_3})^{3/2} \Rightarrow k = \frac{\text{Rate}}{(P_{\text{CH}_3\text{OCH}_3})^{3/2}}$$

Therefore, unit of rate constants (k) = $\frac{\text{bar min}^{-1}}{\text{bar}^{3/2}} = \text{bar}^{-1/2} \text{ min}^{-1}$

Question 3.5:

Mention the factors that affect the rate of a chemical reaction.

Answer 3.5:

The factors that affect the rate of a reaction are as follows.

- > Concentration of reactants (pressure in case of gases)
- > Temperature
- > Presence of a catalyst

Question 3.6:

A reaction is second order with respect to a reactant. How is the rate of reaction affected if the concentration of the reactant is (i) doubled (ii) reduced to half?

Answer 3.6:

Let the concentration of the reactant be $[\text{A}] = a$

Rate of reaction, $R = k[\text{A}]^2 \Rightarrow ka^2$

(i) If the concentration of the reactant is doubled, i.e., $[\text{A}] = 2a$, then the rate of reaction would be

$$R' = k(2a)^2 \Rightarrow 4ka^2 = 4R$$

Therefore, the rate of reaction would be increase by 4 times.

(ii) If the concentration of the reactant is reduced to half, i.e., $[\text{A}] = \frac{1}{2}a$,

Then the rate of the reaction would be $R' = k\left(\frac{1}{2}a\right)^2 = \frac{1}{4}ka = \frac{1}{4}R$

Therefore, the rate of reaction would be reduced to $\frac{1}{4}^{\text{th}}$.

Chemistry 12 3 Section Review Answers

AW Rasmussen



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