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There are a total of 17 pages (40 questions) on the exam. You **MUST** show your work on the exam **AND** turn in this exam along with your ParSCORE form. Circle your answers on the exam.

Carefully bubble in your Perm number, form letter, and answers on your ParSCORE form. Fill in all bubbles completely.

You must show photo ID (preferably your Access Card) when you turn in your exam and ParSCORE form.

1. When hydrocyanic acid (HCN) is added to sodium hydroxide (NaOH), what is the net ionic equation for the reaction that occurs?

- a) $\text{H}^+ + \text{OH}^- \rightleftharpoons \text{H}_2\text{O}$
 b) $\text{HCN} + \text{OH}^- \rightleftharpoons \text{CN}^- + \text{H}_2\text{O}$
 c) $\text{HCN} + \text{Na}^+ + \text{OH}^- \rightleftharpoons \text{H}_2\text{O} + \text{CN}^- + \text{Na}^+$
 d) $\text{HCN} + \text{NaOH} \rightleftharpoons \text{NaCN} + \text{H}_2\text{O}$
 e) $\text{H}^+ + \text{CN}^- + \text{Na}^+ + \text{OH}^- \rightleftharpoons \text{H}_2\text{O} + \text{NaCN}$

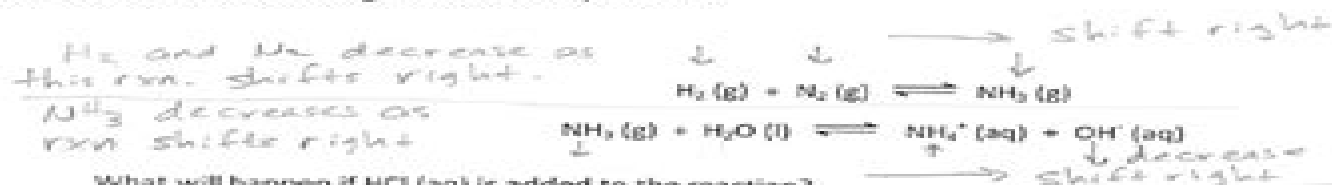
HCN is a weak acid.

Therefore it does not dissociate to a large extent.

Major species: HCN , Na^+ , OH^- , H_2O
 ↑ ↑
 strongest acid strongest base

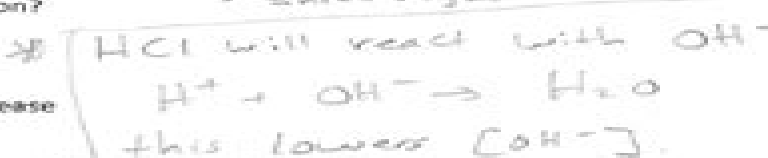


2. Consider the following reactions at equilibrium:



* NH_4^+ increase as rxn. shifts right

- a) The partial pressure of $\text{N}_2(\text{g})$ will decrease
 - b) The partial pressure of $\text{H}_2(\text{g})$ will increase
 - c) The base dissociation constant (K_b) for NH_3 will decrease
 - d) The concentration of $\text{NH}_4^+(\text{aq})$ will decrease
 - e) There will be no change



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Chemistry 1b Practice Final

Ensheng Dong



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