CHEMISTRY I - 1st SEMESTER EXAM REVIEW

STRATEGY: Start by reading through your notes to refresh your memory on these topics. Then, use this review sheet as a starting point to identify the areas on which you need to spend more study time. For those areas, go back to homework assignments, quizzes, and reviews to practice more problems. I would also recommend going through the website links of tutorials and practice quizzes.

FORMAT:

- Questions will be multiple-choice.
- A formula bank will be provided in addition to any values that you might need (electronegativity, etc.), but you will NOT be given "formulas" for items listed in the VOCAB sections (average atomic mass, % error, etc.).

Abc	omic Structure – Ch. 3		500						
1.	 a. Atoms contain negative particles called electrons. b. The mass of an electron is 9.11 × 10⁻²⁸ g. 		 Complete the table for the following isotopes. 						
				Symbol	Zn				
				Atomic #		20			
	 Atoms contain neutral particles called neutral 	roms.		Mass #	65		74	40	
	d. Atoms contain a dense, positive nucleus.			# of protons		1.00	34		
				# of neutrons		21			
2.				# of electrons				1.6	
	billiard ball model to the electron cloud model.			Calculate the average atomic mass of copper if 69.17% or					
3			the copper atoms occurring in nature are 10 Cu and 30.83						
	number, for the following isotopes.			are ⁶⁵ Cu.					
	a. carbon-14 c. nickel-63			VOCAB: isotope average atomic mass					
	b. chromium-53 d. zirconium-92		mole						
Mag	itter - Ch. 1			00			SOCIETA PER	emotest	
6.	Classify the following substances as solid, liquid, gas, or			Classify the folio		chemical	or physics	al changes.	
	plasma based on their properties.			a. cutting wire b. ripening tomato c. apple slices turning brown					
	 a. flexible volume, high KE, particles can disperse freely. 								
	 flexible volume, very high KE, particles are charged. 								
	 fixed volume, very low KE, orderly particles. 			d. compressing a gas					
	d. fixed volume, low KE, particles can move p		 Classify the following properties as physical or chemical. 						
7.			 a. melts at 68 	0°C					
8.				b. corrosive					
	heterogeneous mixture, or solution.		 reacts violently with water 						
	a. graphite (carbon) b. grape juice c. table salt (NaCl) d. pepper			d. decomposes in air e. magnetic					
				law of definite composition law of multiple proportions					
				Mic	easurement - Ch. 2			THE COLUMN	Partie Partie Partie
11.	In a lab, the average measured density for Pre-1982 pennies was 7.98 g/cm ³ . Given that the literature value for the density is 8.92 g/cm ³ , calculate the percent error.			Perform the folk					
				a. 65.2 mm = 3	den				
				b. 2.3 kg = 7g			02 km - 7		
12.	How many sig figs are in the following numbers?			How many milli					
	a. 2.35 c. 89.70			Mrs. J. spent la:					
	b. 34,000 d. 0.0052			spent 5.5 min o		ib. how im	any hours	did it take he	
13.	Convert the following numbers into or out of scientific		to grade all 95 labs?						
	notation.			Calculate the density from the slope of a "Mass vs.					
	a. 548,000 c. 1.200 x 10 ⁻⁹			Volume" graph.					
	 b. 0.0000770 d. 9.25 x 10⁷ 			Record the app			ps when m	reasuring.	
14.	Osmium is the densest element with a density of 22.57			VOCAB: accuracy vs. precision					
	giom ³ . Find the mass of a 56.2 cm ³ sample of o	opmium.							
Ele	ectrons in Atoms - Ch. 4 & 5								
		and a second of the				The second second	Control of the Control		

- 20. Calculate the wavelength if the frequency is 2.5 x 10° Hz.
- 21. Find the energy of a photon if frequency is 7.31 x 10¹⁴ Hz.
- 22. Describe how Bohr's model explains the bright lines (red.
- green, violet, violet) in the emission spectrum of hydrogen.
- 23. What is the primary difference between the modern model of the atom and Bohr's model?
- 24. Draw orbital diagrams for the following elements.

- Symbol Atomic # Orbital Diagram
- Explain why chromium's electron configuration is [Ar] 4s '3d' instead of the expected configuration of [Ar] 4s'3d".

Chemistry 1st Semester Exam Review Answers

Mary Millhollon, Richard H. Langley

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Nursing School Entrance Exam Review William H. Brown, Frederick A. Bettelheim, Jerry March, 2004-04 Artificial Intelligence in Education. Posters and Late Breaking Results, Workshops and Tutorials, Industry and Innovation Tracks, Practitioners, Doctoral Consortium and Blue Sky Andrew M. Olney, Irene-Angelica Chounta, Zitao Liu, Olga C. Santos, Ig Ibert Bittencourt, 2024-07-01 This volume constitutes poster papers and late breaking results presented during the 25th International Conference on Artificial Intelligence in Education AIED 2024 which took place in Recife Brazil during July 8 12 2024 The 18 full papers and 92 short papers were carefully reviewed and selected from 200 submissions They are organized in topical sections as follows Part One Blue Sky Industry Innovation and Practitioner WideAIED and Late Breaking Results Part Two Late Breaking Results Doctoral Consortium Workshops and Tutorials 5 Steps to a 5 AP Chemistry John Moore, 2003-08-22 For the more than one million students taking the AP exams each year Boxed guotes offering advice from students who have aced the exams and from AP teachers and college professors Sample tests that closely simulate real exams Review material based on the contents of the most recent tests Icons highlighting important facts vocabulary and frequently asked questions Websites and links to valuable online test resources along with author e mail addresses for students with follow up questions Authors who are either AP course instructors or exam developers Encyclopedia 2001,2001 Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi,2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching Resources in Education ,1990-05 Study Guide, Introductory Chemistry, a Foundation; Introductory Chemistry; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000 5 Steps to a 5: AP **Chemistry 2020** Mary Millhollon, Richard H. Langley, 2019-08-02 Get ready to ace your AP Chemistry Exam with this easy to follow multi platform study guideThe immensely popular test prep guide has been updated and revised with new material and is now accessible in print online and mobile formats 5 Steps to a 5 AP Chemistry 2020 introduces an easy to follow effective 5 step study plan to help you build the skills knowledge and test taking confidence you need to reach your full potential The book includes hundreds of practice exercises with thorough answer explanations and sample responses You ll learn how to master the multiple choice guestions and achieve a higher score on this demanding exam Because this guide is accessible in print and digital formats you can study online via your mobile device straight from the book or any combination of the three This essential guide reflects the latest course syllabus and includes four full length practice exams 2 in the book and 2 online

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