

Quiz#1 Atomic Structure and the Periodic Table, 1.1-1.5**Total: /28 marks****Part 1: Identify the choice that best completes the statement or answers the question.**

1. The person given credit for developing the first modern periodic table is
 - A. Dalton
 - B. Mendeleev**
 - C. Democritus
 - D. Chadwick
 - E. Thomson
2. Two atoms are isotopes if they have
 - A. different atomic numbers
 - B. the same mass number, but different atomic numbers
 - C. the same number of protons and neutrons
 - D. the same number of electrons, but a different number of neutrons
 - E. the same atomic number, but a different mass number**
3. The element found in the periodic table in Group 6 and Period 4 is
 - A. hafnium, Hf
 - B. gadolinium, Gd
 - C. selenium, Se
 - D. chromium, Cr**
 - E. lithium, Li
4. Which of the following matches of group number and common name is incorrect?
 - A. Group 2 - alkaline earth metals
 - B. Group 17 - halogens
 - C. Group 6 - lanthanides**
 - D. Group 18 - noble gases
 - E. Group 1 - alkali metals
5. Elements with atomic numbers from 57 to 70 are classified as which of the following?
 - A. transuranic elements
 - B. lanthanides**
 - C. noble gases
 - D. actinides
 - E. transition metals
6. Element X is a light, very reactive metal that forms an oxide with the formula X_2O . This element reacts with water to release hydrogen. Identify the element.
 - A. magnesium
 - B. sodium**
 - C. calcium
 - D. copper
 - E. silver

Chemistry 601 Answers

Gaston Vernin, Michel Chanon



Chemistry 601 Answers:

A Textbook of MEDICINAL CHEMISTRY - III (BP601T) Dr. Gopi Reddy Alugubelli, Mr. Sushant Kumar Sharma, Dr. Susmita Basak, Ms. Priti. R. Shirsath, Dr. Arti Gupta, 2024-08-06 The introduction of the book Medicinal Chemistry III makes me incredibly happy This book's content has been painstakingly created to conform to the Pharmacy Council of India's prescribed curriculum for students pursuing a bachelor's degree in pharmacy To make the subject easier for students to understand an attempt has been made to research it using as simple a vocabulary as possible Many images throughout the book including flowcharts and diagrams help students understand difficult concepts The genuine hope of the author is that readers of this book academicians and students alike will find something of value The pharmaceutical product development process serves as the cornerstone for the formulation development process The formulation scientist bears the responsibility of monitoring various material parameters such as API and excipients formulation process parameters dosage forms and other related aspects during the product development process This book provides straightforward and understandable explanations of a wide range of formulation development related subjects including dose I'm hopeful that this book will be well received by both teachers and students We are willing to consider suggestions about any and all facets of the industry Any faults or deviations that may have gone unnoticed are entirely our fault and we would be very grateful if readers could point them out to us if they did *Programmed Problem-solving for First-year Chemistry* Arnold B. Loebel, 1983

Theoretical Chemistry from the Standpoint of Avogadro's Rule & Thermodynamics Walther Nernst, 1923

College Chemistry David Elliott Goldberg, Clyde R. Dillard, 1974 **Answers to questions prescribed by medical state boards** Robert Borneman Ludy, 1907 **Journal of the Chemical Society** Chemical Society (Great Britain), 1883
Titles of chemical papers in British and foreign journals included in Quarterly journal v 1 12 **Computer Aids to Chemistry** Gaston Vernin, Michel Chanon, 1986 Journal - Chemical Society, London Chemical Society (Great Britain), 1883

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The Publishers' Trade List Annual, 1905 *Science* John Michels (Journalist), 1909 Since Jan 1901 the official proceedings and most of the papers of the American Association for the Advancement of Science have been included in

Science *Nature* Sir Norman Lockyer, 1924 **Journal of the Society of Chemical Industry**, 1923 The Chemical Trade Journal and Chemical Engineer, 1909 **Chemical Abstracts**, 1927 *Aquatic Chemical Kinetics* Werner Stumm, 1990-08-09 *Aquatic Chemistry An Introduction Emphasizing Chemical Equilibria in Natural Waters* Second Edition Edited by Werner Stumm and James J Morgan This second edition of the renowned classic unites concepts applications and techniques with the growing amounts of data in the field Expanded treatment is offered on steady state and dynamic models employing mass balance approaches and kinetic information New chapters address such topics as environmental aspects of aquatic chemistry new material on organic compounds in natural water systems the use of stable and radioactive isotopes in chemical and physical processes the latest advances in marine chemistry solid solution interface kinetic considerations of equilibria metal ligand interactions and an expanded compilation of thermodynamic data for important reactions in natural water systems 1981 0 471 04831 3 Cloth 780 pp 0 471 09173 1 Paper *Chemical Processes in Lakes* Edited by Werner Stumm This is a multidisciplinary analysis of recent research on the physical chemical and biological processes in aquatic systems Coverage includes distribution of elements and compounds in water and sediments sedimentation and sediment accumulation of nutrients and pollutants eutrophication and acidification atmospheric deposition redox related geochemistry and sediment water exchange of nutrients and metals sediment dating and paleolimnology and steady state and dynamic models Most chapters focus on the role of biological processes and the coupling of elemental cycles by organisms 1985 0 471 88261 5 435 pp *Principles of Aquatic Chemistry* Francois M M Morel Here is a quantitative treatment of the chemical principles that govern the composition of natural waters Features include an in depth examination of the use of conservation principles in chemical systems a review of thermodynamic and kinetic principles applicable to aquatic systems and a novel presentation of a systematic methodology for equilibrium calculations Detailed coverage is provided on the topic of aquatic chemistry following the traditional divisions of acid base precipitation dissolution coordination redox and surface reactions 1983 0 471 08683 5 446 pp

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