

- C) ii only
- D) i and ii

E) none of the above

8. A solution contains 11% by mass of sodium chloride. This means that

- A) there are 11 g of sodium chloride in in 1.0 mL of this solution
- B) 100 g of the solution contains 11 g of sodium chloride
- C) the density of the solution is 11 g/mL
- D) the molarity of the solution is 11
- E) None of the above

9. Which solution contains the largest number of moles of chloride ions?

- A) 10.0 mL of 0.500M BaCl_2
- B) 4.00 mL of 1.000M NaCl
- C) 7.50 mL of 0.500M FeCl_3
- D) 25.00 mL of 0.400M KCl
- E) 30.00 mL of 0.100M CaCl_2

10. With which of the following will the potassium ion form an insoluble salt?

- A) chloride
- B) sulfate
- C) carbonate
- D) sulfate and carbonate
- E) none of the above

Short Answer:

1. What volume of a 9.00 mol/L hydrobromic acid solution is required to make 4.00 L of a 0.0200 mol/L solution by dilution? (3T, 1C)

2. How many grams of potassium nitrate are required to prepare 4.60×10^2 mL of 0.850 M solution? (3T, 1C)

3. Rubbing alcohol is commonly used as an antiseptic for small cuts. It is sold as a 75% (v/v) solution of isopropyl alcohol in water. What volume of isopropyl alcohol is used to make 400 mL of rubbing alcohol? (3T, 1C)

Chemistry 16 Solutions Answers

J Ma



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