

Calculating Half-life – Answers

As we know, the half-life of a substance is the time taken for half of that substance to decrease to one half of its original amount. This relationship can be shown as:

$$A = A_0 \left(\frac{1}{2}\right)^n \quad \text{where: } A_0 = \text{original activity of the substance;} \\ \text{or } A = A_0 \times (0.5)^n \quad A = \text{current activity} \\ n = \text{number of half-life periods that have elapsed.} \\ \text{also } n = \frac{\text{total time}}{\text{half life}} \quad \text{and} \quad \text{half life} = \frac{\text{total time}}{n}$$

Example:

A sample of Radon-222 has a half-life of 4.20 days. If the original activity was 3.50 MBq, what will the activity be after three half lives

$$A_0 = 3.50 \text{ MBq} \quad A = A_0 \times (0.5)^n \\ n = 3 \quad = 3.5 \times (0.5)^3 \\ A = ? \quad = 3.5 \times 0.125 \quad A = 0.440 \text{ MBq} \\ = 0.4375 \text{ MBq}$$

Activity:

If the original activity of a sample is 21.0 kBq and it has a half-life of 3.00 days, how much will be left after 12.0 days?

$$A_0 = 21 \text{ kBq} \quad A = A_0 \times (0.5)^n \\ n = 12 \div 3 \quad = 21 \times (0.5)^4 \\ = 4 \quad = 21 \times 0.0625 \quad A = 1.31 \text{ kBq} \\ = 1.3125$$

Another way of writing the above equation is shown below. This re-arrangement of the formula can be used when you need to find n .

$$\frac{2^n}{1} = \frac{A_0}{A}$$

Example: Work through this one with your teacher

1. Half-life of Tc-99m is approximately 6 hours.

- a. A sample of Tc-99 has an activity of 8.0×10^3 Bq. How many hours should it take for its activity to drop to 5.0×10^2 Bq?

Half-life = 6 hours

$$A = 5.0 \times 10^2 \text{ Bq} \\ A_0 = 8.0 \times 10^3 \text{ Bq}$$

$$\frac{2^n}{1} = \frac{A_0}{A} = \frac{8000}{500}$$

$$2^n = 16$$

$$n = 4$$

$$\text{total time} = n \times \text{half life} = 4 \times 6 \\ \text{total time} = 24 \text{ hours}$$

- b. If it takes 18 hours for the sample to reach a medical clinic for use, what is the minimum amount that must be transported to the medical clinic if 0.2 g is required for immediate injection?

$$A_0 = ? \\ A = 0.2 \text{ g} \\ \text{Total time} = 18 \text{ hours} \\ \text{Half life} = 6 \text{ hours} \\ n = 18 \div 6 = 3$$

$$A = A_0 \times (0.5)^n \\ 0.2 = A_0 \times (0.5)^3 \\ 0.2 = A_0 \times 0.125 \\ A_0 = 0.2 \div 0.125 \\ = 1.6 \text{ g of sample}$$

Chemistry 1 20worksheet 3 3 Half Life Answers

Arshad Iqbal



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