Omes: $A = (n-1)d^5ns^1$, $B = (n-1)d^1ns^2$, Here n = 4

[D.B.-16]

- a. What is common-ion effect?
- b. In which orbital between 3d and 4p electron will enter first?
- Determine the four quantum number of the last electron of the element A.
- d. Which among the elements A and B will produce colored compound? Explain with reason.

Answer to the question no. 6

The process by which the dissociation rate of the weak electrolyte is reduced due to the presence of a strong electrolyte is known as similar-ion effect.

In which orbital the electron will enter first is determined by Aufbau's principle or (n + I). According to the principle, the orbital of which (n + I) value is low has low energy and electron will enter to it first. But if the (n + I) value of two orbitals are same, then electron will enter the orbital whose n value is lower. For both 3d and 4p orbital the (n + I) value is 5. But for 3d orbital the n value is 3 and for 4p orbital the n value is 4. So, electron will enter 3d orbital first.

The last electron of the given element A enters into (n − 1)d orbital.

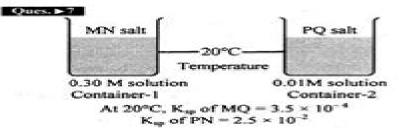
As, the value of n is 4.

So, the value of four quantum numbers of the last electron are :

(i) Principal quantum number = n - 1 = 4 - 1 = 3

(ii) The value of subsidiary quantum number, I for d orbital is 2.

fiii) Magnetic quantum number m = 2



[D.B.-16]

- a. Write down the law of mass action.
- b. Which between H₂SO₃ and HNO₃ is more acidic and why?
- If the MN solution of the stem is saturated then calculate the solubility product of MN at 20°C.
- If the two solutions in the stem are mixed, will there be any ppt? Explain with reason.

Answer to the question no. 7

At a fixed teperature the rate of any reaction is proportional to the active mass (partial pressure in case of liquid and concentration in case of gas) of it's reactants. This is the law of mass action.

Between H₂SO₃ and HNO₃, HNO₃ is more acidic. Because we know that the acid whose oxidation number of the central atom is more is more acidic.

Here for H2SO3 the oxidation number of the central atom S is

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- 1. Understanding the eBook Chemistry 1st Paper Question Hsc 2bing
 - The Rise of Digital Reading Chemistry 1st Paper Question Hsc 2bing
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry 1st Paper Question Hsc 2bing
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 1st Paper Question Hsc 2bing
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry 1st Paper Question Hsc 2bing
 - Personalized Recommendations
 - Chemistry 1st Paper Question Hsc 2bing User Reviews and Ratings
 - Chemistry 1st Paper Question Hsc 2bing and Bestseller Lists
- 5. Accessing Chemistry 1st Paper Question Hsc 2bing Free and Paid eBooks
 - Chemistry 1st Paper Question Hsc 2bing Public Domain eBooks
 - Chemistry 1st Paper Question Hsc 2bing eBook Subscription Services
 - Chemistry 1st Paper Question Hsc 2bing Budget-Friendly Options

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 - Chemistry 1st Paper Question Hsc 2bing Enhanced eBook Features
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 - Highlighting and Note-Taking Chemistry 1st Paper Question Hsc 2bing
 - Interactive Elements Chemistry 1st Paper Question Hsc 2bing
- 8. Staying Engaged with Chemistry 1st Paper Question Hsc 2bing
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 1st Paper Question Hsc 2bing
- 9. Balancing eBooks and Physical Books Chemistry 1st Paper Question Hsc 2bing
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 1st Paper Question Hsc 2bing
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 1st Paper Question Hsc 2bing
 - Setting Reading Goals Chemistry 1st Paper Question Hsc 2bing
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 1st Paper Question Hsc 2bing
 - Fact-Checking eBook Content of Chemistry 1st Paper Question Hsc 2bing
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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