Section 6.1: Organizing the Elements

- 1. Group I on the periodic table is called Alkali Metals
- 2. Group 2 on the periodic table is called Alkaline Earth Metals
- 3. Elements in groups 3-12 are called Transition Metals
- Group 17 on the periodic table is called Halogens
- 5. Group 18 on the periodic table is called Noble Gases
- Each period on the periodic table corresponds to a principal energy level
- 7. The 2 series that make up the inner transition metals are Lanthanides and Actinides
- Complete the following chart:

| Element | Symbol | Group # | Period # | Group Name | Metal, Metalloid, Nonmetal |
|-----------|--------|---------|----------|-----------------------|----------------------------|
| Zirconium | Zr | - 4 | 5 | Transition Metals | Metal |
| Radium | Ra | 2 | 7 | Alkaline Earth Metals | Metal |
| Bismuth | Bi | 1.5 | 6 | Nitrogen Group | Metal |
| Gallium | Ga | 1.3 | 4 | Boron Group | Metal |

- 9. What are triads? Sets of 3 elements with similar properties Whose idea were they? Dobereiner
- 10. The modern periodic table is arranged by increasing atomic number. Whose idea was this? Moseley
- Mendeleev developed the first periodic table, which was arrange by increasing atomic mass
- 12. Determine which of the following pairs will have similar properties (Yes or no):
 - a. Ca and Be? Yes same group (Group 2 Alkaline Earth Metals)
 - b. Sr and Y? No not the same group
 - c. C and Ge? Yes same group (Group 14 Carbon Group)

Section 6.3: Periodic Trends

- As you go across a period, the atomic radius (size) decreases because the attraction between protons
 and electrons increases, which makes the atom smaller (opposite of tug of war)
- As you go down a group, the atomic radius (size) increases because the number of principal energy levels increases
- Circle the atom with the largest atomic radius and put a square around the atom with the smallest radius:
 - a. (Li) Ne F B
 b. N Bi S P
 c. Na (Sr) Al (Cl
- 4. Ionization energy is the energy required to remove an electron.

<u>Chemistry 6 Organizing The Elements Section Review</u> <u>Answers</u>

Arshad Iqbal

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