

**Experiment 1 - Volumetric Analysis (Unknown  $\text{Al}^{3+}$  Concentration)**

- Given an aqueous solution of  $\text{p}(\text{HCl})_3$  of unknown concentration
- Goal is determine the concentration of  $\text{Al}^{3+}$
- Chemistry used: solubility
- Method: convert  $\text{Al}^{3+}$  (soluble, clear) to  $\text{Al}(\text{OH})_3$  (insoluble, yellow)
- pH
  - Low pH ( $\text{pH} \approx 1$ ) test across  $\text{H}^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+$
  - High pH correct protein stain  $\rightarrow \text{H}_2\text{O}$  test *insoluble*
  - Very high pH:  $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$
  - Neutral pH:  $\text{H}_2\text{O}$  (soluble)
  - pH 5: buffer + ammonium acetate ( $\text{NH}_4^+/\text{NH}_3$ )
- Calculations:
  - (mass of reactants + precipitate) - (mass of reactants) = mass of precipitate
  - (mass of precipitate)/(molar mass)  $\rightarrow$  mol of  $\text{Al}(\text{OH})_3$  = mol of  $\text{Al}^{3+}$
  - mol of  $\text{Al}^{3+}$ /volume = concentration of  $\text{Al}^{3+}$
- Systematic errors (interferences): "buffer" process, pH control, equilibrium constant

**Experiment 2 - Volumetric Analysis (Unknown Acid Identity)**

- Given an unknown acid
- Goal is determine the identity of the unknown acid (monoprotic)
- Chemistry used: acid base rxn
- Method:
  - (1) standardised NaOH ( $\text{NaOH} \rightarrow \text{NaCl}$ )
    - NaOH concentration is tested by base with alcohol- $\text{CO}_2$  & benzoic acid
    - in  $\text{NaCl}$  is better than  $\text{NaOH}$  by it has more  $\text{H}_2\text{O}$  (rough and a little bit more  $\rightarrow$  more  $\text{H}_2\text{O}$ )
    - Indicator = phenolphthalein
  - (2) titration  $\rightarrow$  water stain
  - (3) pH meter  $\rightarrow$  pH
    - pH 4  $\rightarrow \text{H}_2\text{O}/\text{H}_2\text{O} \rightarrow$  after chemical process
    - $\text{H}^+$  is an important  $\text{H}_2\text{O}$   $\rightarrow$  correct
- Calculations:
  - (volume of NaOH)/(molar mass of NaOH)  $\rightarrow$  moles of NaOH = moles of unknown
  - (mass of unknown)/(moles of unknown) = molar mass
  - $\text{H}^+$  then also  $\rightarrow$  greatest rate of change = eq point
- Acid dissociation constant ( $K_a$ )
  - $\text{HA} + \text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{A}^-$
  - $K_a = [\text{H}^+][\text{A}^-]/[\text{HA}]$
- Henderson-Hasselbalch equation  $\text{pH}_a = \text{pH} - \log([\text{A}^-]/[\text{HA}])$ 
  - At 0.5 eq point:  $[\text{A}^-]/[\text{HA}] = 1/1$ , and  $\log(1) = 0$  so  $\text{pH}_a = \text{pH}$
  - At 1 eq point:  $[\text{A}^-]/[\text{HA}] = 1/0$  and  $\log(0) = -\infty$  so  $\text{pH}_a = \text{pH} + \infty$
  - At 0 eq point:  $[\text{A}^-]/[\text{HA}] = 0/1$ , and  $\log(0) = -\infty$  so  $\text{pH}_a = \text{pH} - \infty$

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