

# Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = ?$$



$$\Delta H^\circ = +184.6 \text{ kJ}$$



$$\Delta H^\circ = -314.4 \text{ kJ}$$



$$\Delta H^\circ = -92.2 \text{ kJ}$$

$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol,  $\Delta H_{\text{comb}}$



$$\Delta H = -277.7 \text{ kJ}$$



$$\Delta H = -393.5 \text{ kJ}$$



$$\Delta H = -285.8 \text{ kJ}$$

$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone,  $\text{CaCO}_3$ , to lime,  $\text{CaO}$ , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{ CaCO}_3(\text{s}) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CaO}(\text{s}) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2(\text{g}) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4.  $\text{CH}_4(\text{g}) + 3 \text{Cl}_2(\text{g}) \rightarrow \text{CHCl}_3(\text{g}) + 3 \text{HCl}(\text{g}) \quad \Delta H = ?$



$$\Delta H^\circ = 92.3 \text{ kJ}$$



$$\Delta H^\circ = -74.8 \text{ kJ}$$



$$\Delta H^\circ = -103.1 \text{ kJ}$$

$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide,  $\text{CH}_2\text{N}_2$ , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = -741.4 \text{ kJ/mol}$$



$$\Delta H = -393.5 \text{ kJ/mol}$$



$$\Delta H = -285.8 \text{ kJ/mol}$$

$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation:  $\Delta H^\circ_f \text{ NO}(\text{g}) = 90.3 \text{ kJ/mol}$  and  $\Delta H^\circ_f \text{ H}_2\text{O}(\text{g}) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction:  $2 \text{Al}(\text{s}) + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow \text{Al}_2\text{O}_3(\text{s}) + 2 \text{Fe}(\text{s}) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess  $\text{Fe}_2\text{O}_3$ , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

# Chem Ap 2013 Response Answers

**Michael Rajnik**



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