

Observation Table

S.No	Burette readings		net vol ⁿ used
	initial	final	
(1)	0.0	14.4	14.4
(2)	14.4	29	14.6
(3)	29	43.6	14.6

Calculation

$$N_1 V_1 = N_2 V_2$$

(acid) (alk)

$$\frac{0.1 \times 10}{10} = N_2$$

$$N_2 = \frac{0.1 \times 14.6}{10} = 0.146 \text{ g/L}$$

$$\text{Strength} = N_{\text{NaOH}} \rightarrow 5.84 \text{ g/L}$$

total g/L

$$\text{strength of NaOH} = (8 - 1) = 7$$

$$8 - 5.84 = 2.16 \text{ g/L}$$

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Experiment no.1

Aim: To find out the percentage composition of the given mixture of NaOH and NaCl.

Apparatus required: Burette, pipette, chemical flask, beaker.

Chemicals required: 0.1 N NaCl, NaOH, NaOH, NaOH, Phenolphthalein, indicator.

End point: Pink to colourless.

Reaction involved:



Procedure

1) Rinse the bottle with the given NaOH solⁿ and standard found immediately.

2) Take 10 ml solⁿ in the burette and record the initial readings.

3) Pipette out the 10 ml of the solution in the chemical flask.

4) Add a drop of indicator in the solution.

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