

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING

Concentration of Solutions

There are three principal ways to express solution concentration in chemistry — percentage by mass, molarity, and molality.

The following table compares these three ways of stating solution concentration. Examining the method of preparation of the three types may help you understand the differences among them.

	Symbol	Meaning	How to prepare
Percentage	%	Grams solute per 100 g of solution	5%: Dissolve 5 g of solute in 95 g solvent.
Molarity	M	Moles solute per liter of solution	5 M: Dissolve 5 mol of solute in solvent and add solvent to make 1 L of solution.
Molality	<i>m</i>	Moles solute per kilogram of solvent	5 <i>m</i> : Dissolve 5 mol of solute in 1 kg of solvent.

PERCENTAGE CONCENTRATION

You will find percentages of solutes stated on the labels of many commercial products, such as household cleaners, liquid pesticide solutions, and shampoos. If your sink becomes clogged, you might buy a bottle of drain opener whose label states that it is a 2.4% sodium hydroxide solution. This means that the bottle contains 2.4 g of NaOH for every 100 g of solution.

Computing percentage concentration is very much like computing percentage composition (see Chapter 6). Both involve finding the percentage of a single component of a multicomponent system. In each type of percentage calculation, the mass of the important component (in percentage concentration, the solute) is divided by the total mass of the system and multiplied by 100 to yield a percentage. In percentage concentration, the solute is the important component, and the total mass of the system is the mass of the solute plus the mass of the solvent.

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