Name: Class: Five Types of Chemical Reaction Worksheet Answers Balance the following reactions and indicate which of the six types of chemical reaction are being represented: 2 NaBr + 1 Ca(OH)₂ → 1 CaBr₂ + 2 NaOH Type of reaction: double displacement 2 NH₂+ 1 H₂SO₄ → 1 (NH₄)₂SO₄ Di) Type of reaction: synthesis c): 4 C₂H₂O + 27 O₂ → 20 CO₂ + 18 H₂O Type of reaction: combustion d): 3 Pb + 2 H₂PO₄ → 3 H₂ + 1 Pb₃(PO₄)₂ Type of reaction: single displacement (0.0) 1 Li₂N + 3 NH₄NO₂ → 3 LiNO₂ + 1 (NH₄)₂N Type of reaction: double displacement Section 3: Predicting the products of chemical reactions 33 2 Ag + 1 CuSO₄ → 1 Ag₂SO₄ + 1 Cu Type: Single Displacement 20 2 Nal + 1 CaCl₂ → 2 NaCl + 1 Cal₂ Type: Double Displacement 300 1 O2 + 1 H2 -> 2 H2O Type: Synthesis Type: Double Displacement 4) 2 AgNO₂ + 1 BaSO₄ → 1 Ag₂SO₄ + 1 Ba(NO₂)₂ 5) 2 HCN + 1 CuSO₄ → 1 H₂SO₄ + 1 Cu(CN)₂ Type: Double Displacement 6) 1 H₂O + 1 Agl → 1 HI + 1 AgOH Type: Double Displacement 7% 4 LiBr + 1 Co(SO₁)₂ → 2 Li₂SO₃ + 1 CoBr₄ Type: Double Displacement 8) 1 LiNO₃ + 1 Ag → 1 AgNO₃ + 1 Li Type: Single Displacement 91 1 N₂ + 2 O₂ → 2 MO₂ Type: Synthesis 100 1 H₂CO₂ → 1 CO₂ + 1 H₂O Type: Decomposition Type: Single Displacement 113 1 AlCl₃ + 3 Cs → 3 CsCl + 1 Al 121 1 Al(NO₁)₁ + 1 Ga → 1 Ga(NO₁)₁ + 1 Al Type: Single Displacement 130 1 CH₂COOH + 2 O₂ → 2 CO₂ + 2 H₂O Type: Combustion 3.40 1 C₀H₀ + 6 O₂ → 4 CO₂ + 4 H₂O Type: Combustion 15) 2 KCI + 1 Mg(OH)₂ → 2 KOH + 1 MgCl₂ Type: Double Displacement 1 Zn + 1 Au(NO₂)₂ → 1 Zn(NO₂)₂ + 1 Au 1.63 Type: Single Displacement 1.77 1 BaS + 1 PtCl₂ → 1 BaCl₂ + 1 PtS Type: Double Displacement

Type: Decomposition

18)

2 Na₂O → 4 Na + 1 O₂

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