

Intermolecular Forces Worksheet

Answers are on page 3 & 4. Do the problems on your own **BEFORE** looking at the answers.

1. Predict the molecular shape of each of the following:

- a. H_2S _____
- b. CCl_4 _____
- c. SO_2 _____
- d. BrF _____
- d. PCl_5 _____

2. List **all** types of IMFs that would occur in each of the following (you should have a good enough understanding of electronegativities to answer all of these, except maybe SO_2 , without look up the electronegativity numbers).

- a. CH_3CF_3 _____
- b. CCl_4 _____
- c. SO_2 _____
- d. BrF _____
- e. $(\text{CH}_3)_3\text{N}$ _____
- f. PCl_5 _____

3. H_2S , O_2 and CH_3OH all have comparable molecular masses. List the **dominant** type of IMF for the pure substances, then rank the strength of each compound based on IMFs within the samples.

(1 = strongest, 2 = in between, 3 = weakest).

Substance	IMF	Relative Strength
HBr		
O_2		
CH_3OH		

Chem 150 Activity On Intermolecular Forces Answers

Michael Seilmaier



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