

## Section 3: Chemistry (continued) - acids & alkaline substances

### (c) Substances - basic substances

These substances are substances which react with acids and therefore it is the mass of particles which contain the basic substance. The number of particles is measured by the coefficient in front of each substance in the reaction.

- a. How many moles of  $\text{NaOH}$  will be formed from 10 tonnes of  $\text{Na}_2\text{CO}_3$  (assuming 100% yield) according to the equation below?



- b. How many moles of  $\text{Ca}_3(\text{PO}_4)_2$  are produced from 10 tonnes of  $\text{CaO}$  (assuming 100% yield) according to the equation below?



- c. How many moles of  $\text{HCl}$  will be produced from 10 tonnes of  $\text{NaCl}$  (assuming 100% yield) according to the equation below?



### (d) Substances - acid substances

#### Substances - acid substances

- Convert grams to moles by dividing it by the molar mass
- Use the molar mass to find the number of moles of acid
- Convert the number of moles of acid to grams by multiplying it by the molar mass

- a. How many grams of  $\text{HCl}$  will be formed from 10 tonnes of  $\text{NaCl}$  (assuming 100% yield) according to the equation below?



- 10 tonnes of  $\text{NaCl}$  is 10 tonnes of  $\text{NaCl}$  (assuming 100% yield) is then molar mass of  $\text{NaCl}$

$$\text{molar mass of NaCl} = \frac{\text{molar mass}}{\text{molar mass}} = 10 \text{ tonnes HCl}$$

# Chem Ch9 Stoichiometry Key

**Ralph A. Burns**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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