

Ionic Bonding

Name: _____
Hour: _____

Information: Polyatomic Ions

The word, "polyatomic" means "many atoms". A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion, SO_4^{2-} . Sulfate is composed of one sulfur atom and four oxygen atoms and overall sulfate has a negative two charge.

Some polyatomic ions:

Sulfate: SO_4^{2-} Cyanide: CN^- Acetate: $\text{C}_2\text{H}_3\text{O}_2^-$ Phosphate: PO_4^{3-} Ammonium: NH_4^+ Hydroxide: OH^- Nitrate: NO_3^- Chlorate: ClO_3^- Carbonate: CO_3^{2-}

Critical Thinking Questions

1. What do all of the polyatomic ions that have the suffix "-ate" have in common?
2. Which two atoms do you think compose the polyatomic ion called "silicate"?
3. What is the difference between calcium nitride and calcium nitrate?

Information: Writing Formulas With Polyatomic Ions

First of all, you must remember that you can never change the formula for a polyatomic ion. Sulfate is always SO_4^{2-} and never $\text{S}_2\text{O}_6^{4-}$ or something else. Following are some examples of chemical formulas that contain polyatomic ions.

Ammonium chloride is formed from one ammonium ion (NH_4^+) and one chloride ion (Cl^-) to give the formula: NH_4Cl . Sodium sulfate requires two sodium ions (Na^+) because sulfate (SO_4^{2-}) has a negative two charge; the formula is: Na_2SO_4 .

Consider calcium hydroxide. Calcium has a positive two charge (Ca^{2+}) and hydroxide has a negative one charge (OH^-). We need two hydroxide ions to combine with one calcium ion so that the overall charge ends up being zero. We write calcium hydroxide like $\text{Ca}(\text{OH})_2$.

Following are some more examples:

potassium acetate: $\text{KC}_2\text{H}_3\text{O}_2$ barium phosphate: $\text{Ba}_3(\text{PO}_4)_2$ magnesium nitrate: $\text{Mg}(\text{NO}_3)_2$ calcium carbonate: CaCO_3

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Minjie Lin



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Encyclopedia of Chemical Processing (Online) Sunggyu Lee, 2005-11-01 This second edition Encyclopedia supplies nearly 350 gold standard articles on the methods practices products and standards influencing the chemical industries It offers expertly written articles on technologies at the forefront of the field to maximize and enhance the research and production phases of current and emerging chemical manufacturing practices and techniques This collecting of information is of vital interest to chemical polymer electrical mechanical and civil engineers as well as chemists and chemical researchers A complete reconceptualization of the classic reference series the Encyclopedia of Chemical Processing and Design whose first volume published in 1976 this resource offers extensive A Z treatment of the subject in five simultaneously published volumes with comprehensive indexing of all five volumes in the back matter of each tome It includes material on the design of key unit operations involved with chemical processes the design unit operation and integration of reactors and separation systems process system peripherals such as pumps valves and controllers analytical techniques and equipment and pilot plant design and scale up criteria This reference contains well researched sections on automation equipment design and simulation reliability and maintenance separations technologies and energy and environmental issues Authoritative contributions cover chemical processing equipment engineered systems and laboratory apparatus currently utilized in the field It also presents expert overviews on key engineering science topics in property predictions measurements and analysis novel materials and devices and emerging chemical fields ALSO AVAILABLE ONLINE This Taylor E mail e reference taylorandfrancis com International Tel 44 0 20 7017 6062 E mail online sales tandf co uk

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