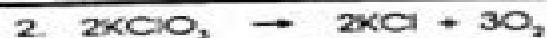


**STOICHIOMETRY:
MOLE-MOLE PROBLEMS**

How many moles of hydrogen are needed to completely react with two moles of nitrogen?

$$\frac{2 \text{ mol N}_2}{1 \text{ mol N}_2} \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} = 6$$

6 moles H₂

How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

$$6 \text{ moles KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = \frac{18}{2} = 9$$

9 moles of O₂

How many moles of hydrogen are produced from the reaction of three moles of zinc with an excess of hydrochloric acid?

$$3 \text{ mol Zn} \times \frac{1 \text{ mol H}_2}{1 \text{ mol Zn}} = 3$$

3 moles of H₂How many moles of oxygen are necessary to react completely with four moles of propane (C₃H₈)?

$$4 \text{ mol C}_3\text{H}_8 \times \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} = 20$$

20 mol of O₂

How many moles of potassium nitrate are produced when two moles of potassium phosphate react with two moles of aluminum nitrate?

$$\frac{2 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} = 6$$

6 moles of KNO₃

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J Ma



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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

The Big Book of Chemistry Teacher Stories Jeff Lark, Stories from years of teaching high school chemistry

Chemistry James N. Spencer, George M. Bodner, Lyman H. Rickard, 2010-12-28 CHEMISTRY

Introduction to Chemistry, Laboratory Manual T. R. Dickson, 1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students

Basic Chemistry William S. Seese, Guido H. Daub, 1985

Resources in Education , 1980

Journal of Chemical Education , 1929

Caveman Chemistry Kevin M. Dunn, 2003 Half a million years ago our ancestors learned to make fire from scratch They crafted intricate tools from stone and brewed

mind altering elixirs from honey Their descendants transformed clay into pottery wool into clothing and ashes into cleansers In ceramic crucibles they won metal from rock the metals lead to colored glazes and glass Buildings of brick and mortar enshrined books of parchment and paper Kings and queens demanded ever more colorful clothing and accessories in order to out class clod hoppers and call girls Kingdoms rose and fell by the power of saltpeter sulfur and charcoal And the demands of everyday folk for glass and paper and soap stimulated the first round of chemical industrialization From sulfuric acid to sodium carbonate From aniline dyes to analgesic drugs From blasting powder to fertilizers and plastics In a phrase From Caveman to Chemist Your guides on this journey are the four alchemical elements Fire Earth Air and Water These archetypal characters deliver first hand accounts of the births of their respective technologies The spirit of Fire for example was born in the first creature to cultivate the flame This spirit passed from one person to another from one generation to another from one millennium to another arriving at last in the pages of this book The spirit of Earth taught folks to make tools of stone the spirit of Air imparted knowledge of units and the spirit of Water began with the invention of spirits Having traveled the world from age to age who can say where they will find their next home Perhaps they will find one in you

Chemistry & Chemical Reactivity John C. Kotz,Paul Treichel,Patrick A. Harman,2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon **Chemical Engineering Library Book List** American Institute of Chemical Engineers.

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