

**CHEM 108**  
**ChemActivity 16: Stoichiometry of Chemical Equations**  
**Solution Key**

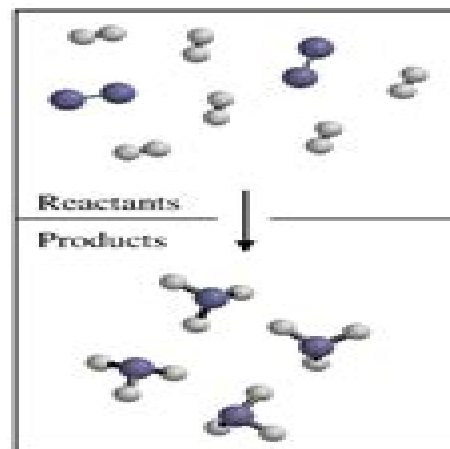
**The Model**

The reaction between elemental hydrogen and elemental nitrogen to produce ammonia is given by eq 1.



The chemical equation can be interpreted in two different ways:

- (1) Microscopically – Three molecules of hydrogen react with one molecule of nitrogen to yield two molecules of ammonia.
- (2) Macroscopically – For every three moles of hydrogen that react, one mole of nitrogen reacts to produce two moles of ammonia.



**Key Questions**

1. While describing the reaction on a “microscopic scale”, why would it be inappropriate to use the following?



On a microscopic scale, it would be “three-halves of a hydrogen molecule react with half a nitrogen molecule to yield one ammonia molecule.” But there is no way to have half of a molecule! (Monatomic N doesn’t react the same as N<sub>2</sub>.)

- 2a. In the synthesis of ammonia from its elements, how many moles of H<sub>2</sub> are consumed for every mole of N<sub>2</sub> that is consumed?

In eq 1, the stoichiometric coefficients in front of N<sub>2</sub> and H<sub>2</sub> are 1 and 3, respectively. So 3 moles of H<sub>2</sub> will be consumed for every 1 mole of N<sub>2</sub> that is consumed.

# Chemactivity 41 Key

**R Pring**



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