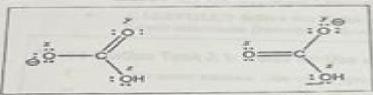
ChemActivity 2: Resonance Structures

(Why is baking sods a much weaker base than hydroxide, though both have an O with a -1 charge?)

Model 1: Two Lewis Structures for Baking Soda (bicarbonate)

There are two different but equally good ways to draw a legitimate Lowis structure of bicarbonate ion.



Neither structure by itself tells the whole story.

To get the most accurate picture of bicarbonate ion you must gverage the two different structures.

Figure 2.1: Bicarbonate Ion (HCO₃ [⊖])

To facilitate the following discussion, the three oxygens of bicarbonate were labeled O., O., and O.,

Construct Your Understanding Questions (to do in class)

A student looks at Model 1 and says she expects the bond between C and O, to be halfway between
a single bond and a double bond (i.e., shorter and stronger than a single bond, but not as short and
strong as a double bond) and that the same will be true of the bond between C and O,.

Use the information in Model 1 to construct an explanation for these conclusions.

- Based on the most accurate picture of bicarbonate ion, what is the charge on O₁? ... O₂? ... O₃?
 (Note that it may be a fraction.) Briefly explain how you calculated your answers.
- Organic chemists say that "the -1 charge on bicarbonate ion is 'delocalized' between O_x and O_y."
 Come up with at least one synonym for the word "delocalized" in the preceding sentence.
- Which do you expect to have a more intense and concentrated "hot spot" of negative charge: methoxide ion or bicarbonate ion? Explain.

CH3 0:0

One definition of a strong base is: "A molecule with a -1 formal charge localized on a single H, and N, or O atom." Construct an explanation for why bicarbonate ion is NOT a strong base.

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