

**K<sub>eq</sub> for FeSCN<sup>2+</sup> with Microlab****INTRODUCTION**

Many chemical reactions, such as acid-base neutralization, are said to "go to completion". Reactants disappear and products are formed, theoretically at a 100% yield. Other reactions reach equilibrium so that there are constant amounts of reactants and products in the mixture. The extent of the reaction is reflected in the value of the equilibrium constant, K<sub>eq</sub>.

For the reaction



the equilibrium constant expression is:

$$K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

Note that the product concentrations are in the numerator, so when more products than reactants are present at equilibrium, the K<sub>eq</sub> value is large. Also note that the equilibrium expression can be written directly from the balanced chemical equation (unlike the rate law). The coefficients in the chemical equation become the exponents in the equilibrium expression.

The equilibrium reaction investigated in this lab experiment is:



The complex ion, FeSCN<sup>2+</sup>, is a blood-red color and strongly absorbs light with a wavelength of 447 nm. The amount of absorbance is directly proportional to the concentration of the complex ion. The greater the concentration, the greater the absorbance. If no FeSCN<sup>2+</sup> is present, the absorbance is zero. By using a very large amount of Fe<sup>3+</sup> and an extremely small amount of SCN<sup>-</sup>, the equilibrium can be pushed so far to the right that we can assume that all the SCN<sup>-</sup> has been converted to FeSCN<sup>2+</sup>. By measuring the absorbance of several mixtures with different concentrations of SCN<sup>-</sup>, we can construct a calibration curve that relates absorbance to the concentration of FeSCN<sup>2+</sup>. Then, in part B, we will use a much smaller concentration of Fe<sup>3+</sup> to create three equilibrium mixtures with a measureable amount of reactants and products.

In part C we will construct the calibration curve and use it to measure the concentration of the complex ion in the equilibrium mixtures. Since initial concentrations of Fe<sup>3+</sup> and SCN<sup>-</sup> can be calculated and the amount of change will be equal to the amount of complex ion that is formed, the equilibrium amounts of each substance can be determined with the help of an ICE table. The equilibrium constant, K<sub>eq</sub>, will be calculated for each equilibrium mixture. To get good precision, all volumes must be very, very carefully measured.

**PART A. PREPARATION OF THE STANDARD CURVE**

Four standard solutions will be prepared by each pair, using the following procedure:

1. Set up 4 large test tubes and label them 1, 2, 3, and 4.

# Chemfax Determinatin Of Keq Prelab Answers

**Daniel F McAuley**



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