Prelab Activity:

Chemical Kinetics

- A student performed experiments 1 and 3 from the table on page 2 of the lab handout. In experiment 1, the reaction mixture took 161 s to turn purple, while in experiment 3, the reaction mixture took 83 s to turn purple.
 - a) Using Equation 5 in the lab handout, determine the rate of reaction for experiments 1 and 3.

b) Using Equation 2, determine the value of the exponent, y (the order of the reaction with respect to I'). Report your value to the nearest whole number. Hint: Remember the total volume of the reaction mixture is 15 mL, so you will need to perform dilution calculations for the [I'] in reactions 1 and 3.

2) Consider the following chemical reactions performed in the kinetics lab:

$$H_2O_2 + 2 \Gamma + 2H_3O^* \rightarrow 4 H_2O + I_2$$

 $2S_2O_3^{2-} + I_2 \rightarrow 2 \Gamma + S_4O_6^{2-}$

Assuming that you followed the procedure correctly, a color change will occur at approximately what moment?

- (a) When all hydrogen peroxide has been converted to water
- (b) When all of the S2O32 has been used up
- (c) When the H3O has been fully neutralized by the acetate buffer
- (d) When the HyO' interacts with the molybdenum catalyst

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Keith James Laidler

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An Examples Course in Reaction Kinetics Ian M. Campbell,1980 Rates and Mechanisms of Chemical Reactions W. C. Gardiner (Jr.), William Cecil Gardiner,1969 Basic Reaction Kinetics and Mechanisms Harold Eric Avery,1974 Theories of Chemical Reaction Rates Keith James Laidler,1969 Chemical Kinetics Noel Roderick Waite,1971

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