

# Pre-Lab Questions (Use a separate sheet of paper to answer the following questions.)

1. In Part 1 of this laboratory, 25.0 mL of a 0.100 M solution of the weak acid acetic acid,  $\text{CH}_3\text{COOH}$ , is titrated with a 0.100 M solution of the strong base sodium hydroxide,  $\text{NaOH}$ .  $K_a$  of acetic acid is  $1.8 \times 10^{-5}$ .

a. Calculate the pH of the equivalence point. Enter this value in the Part 1 Data Table.

b. Five indicators, along with their  $pK_a$  values and their color change ranges are shown below.

$5.6 \times 10^{-10} \approx \frac{x^2}{0.050}$   $x = 5.3 \times 10^{-6}$   $pOH = 5.28$   $pH = 8.72$

	$pK_a$	1	2	3	4	5	6	7	8	9	10	11	12	13
Methyl Orange	3.40		Red											
Methyl Red	4.95		Red											
Phenolphthalein	9.4					Colorless							Red	
Bromothymol Blue	7.1				Yellow							Blue		
Thymolphthalein	10.0													

$\text{CH}_3\text{COO}^- + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COOH} + \text{OH}^-$

I Based on the pH calculated in 1a., select the appropriate indicator for the weak acid-strong base titration. Enter the selection in the Part 1 Data Table.

C 2. In Part 2, 25.0 mL of a 0.10 M solution of the weak base ammonia,  $\text{NH}_3$ , is titrated with a 0.10 M solution of the strong acid hydrochloric acid,  $\text{HCl}$ .  $K_b$  of ammonia is  $1.8 \times 10^{-5}$ .

E a. Calculate the pH of the equivalence point. Enter this value in the Part 2 Data Table.

b. Select an indicator for the weak base-strong acid titration. Enter the selection in the Part 2 Data Table.

# Chemfax Lab Answers

**L Cohen**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'L Cohen' text.

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