

100

27. It is necessary to know the formula of a compound in order to calculate its percentage composition.

AT

28. If the percent by mass of carbon in methane, CH_4 , is 75%, then 100 grams of methane contains 75 g of percent of hydrogen.

AT

29. The formula for methane, CH_4 , is both a molecular and an empirical formula.

100

30. The empirical formula of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, is CH_2O .

C. Identification

Write the following particles in the space provided. Show your work.

27. How many neutrons of 23 in 5.1, at 10^{22} atoms of 23?

28. How many representative particles are in 2.18, at 10^{22} and MgCl_2 ?

29. What is the gram molecular mass of methane, CH_4 ?

30. Find the mass of each of the following:

a. 2.43, at 10^{22} and H_2O_2

b. 2.43 and H_2O_2

c. 4.86 and 1.36

d. 4.86 and 4.3

Chemfiesta Chemical Quantities Answer Key

Sebastian Brünink



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