

Observation of a Limiting Reagent Lab

NAME: _____ DATE: _____ PERIOD: _____

Background:

A balanced equation shows the amount of reactants needed to make a certain amount of products. A reaction with two reactants can have one reactant in excess amount and the other reactant in limited amounts. The amount of one reactant used will be based on the amount of the other reactant present. For example, water can be produced by the following equation $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$.

If you have 20 liters of oxygen gas at STP then you would need 40 liters of hydrogen gas. If you have less than the 40 liters needed then the hydrogen is the limiting reagent and the oxygen is the excess reagent. If you have more than 40 liters then the hydrogen is the excess reagent and the oxygen is the limiting reagent. In this lab the limiting reagent and excess reagent will be found by a reaction between a strip of magnesium and hydrochloric acid.

Procedure:

- 1) Obtain the mass of a medium size clean dry test tube.
- 2) Place a 1-inch strip of magnesium ribbon into the test tube. Then obtain the mass of the test tube and **magnesium**.
- 3) Place the test tube into a test tube rack and add 10.0 milliliters of **3.0 M** (←**this is the molarity**) **hydrochloric acid** into the test tube. Observe.
- 4) When the reaction is over dump the test tube contents into the central waste beaker. Rinse the test tube and return to the rack. Rinse the graduated cylinder and leave by the test tube rack.

Data:

Mass of test tube	grams
Mass of test tube and magnesium	grams
Mass of magnesium	grams
Volume of acid used	milliliters
Molarity of acid used	Molar

1) Write a balanced equation for the reaction. (**Not** reaction from background!!!!!!)
Hint: Identify type of reaction by the reactants then predict products.)

- 2) Using the balanced equation, draw the test tube during the reaction. Be very specific and label all four species from the balanced equation.
- 3) Draw the test tube at the end of the reaction. Be very specific and label all species.

Chemfiesta Limiting Reactant Lab Answers

Lifelique



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Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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