

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

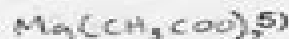
$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left( \frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW<sub>Ca(NO<sub>2</sub>)<sub>2</sub></sub> = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW<sub>NaOH</sub> = 39.997 g/mol

$$n = 98 \text{ g} \left( \frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

MW<sub>HCl</sub> = 36.461 g/mol

$$n = (1.2 \text{ g}) \left( \frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW<sub>NH<sub>3</sub></sub> = 17.031 g/mol

$$n = 45 \text{ g} \left( \frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

# Chemfiesta Of Stiochiometric Calculation Answers

**KA. Gavhane**



## **Chemfiesta Of Stiochiometric Calculation Answers:**

**You Can Do Chemistry** Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

**Stoichiometry and Process Calculations** Mr. Rohit Manglik, 2024-01-22 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels

**Stoichiometric Calculations** Lifeliqe, 2019 This lesson plan covers how to calculate moles and volume ratios for balanced chemical equations

*Bite-Sized Chemistry Calculations* John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance

CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices

**Introduction to Process**

**Calculations Stoichiometry** KA. Gavhane, 2012

**Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied

objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

Stoichiometric Calculations: Known Problems and Proposed Solution at a Chemistry-mathematics Interface  
Werner Dierks,1981

Factors in the Performance of Stoichiometric Calculations in Chemistry Ronald Zachariah

Fernandez,1990 Students' Conceptions of Stoichiometry at the Submicro Level Sin-Yan Chan,□□□,2017-01-27 This

dissertation Students Conceptions of Stoichiometry at the Submicro Level by Sin yan Chan was obtained from The University of Hong Kong Pokfulam Hong Kong and is being sold pursuant to Creative Commons Attribution 3 0 Hong Kong License The content of this dissertation has not been altered in any way We have altered the formatting in order to facilitate the ease of printing and reading of the dissertation All rights not granted by the above license are retained by the author Abstract Stoichiometry is an important topic in chemistry It tells how many reactants are required to produce a certain amount of product in terms of mass mole and volume Learning stoichiometric calculation involves the understanding of certain concepts such as the mole stoichiometric ratios and chemical equations Some studies attributed the failure in learning stoichiometry to the unfamiliarity with the amount of substance in a mole and students weakness in the mathematical ability Nevertheless how students connect the submicro level and the symbolic level in learning stoichiometry was not discussed widely In this study two examples of chemical reactions with different levels of difficulties were used to probe students conceptual understanding in stoichiometry at the submicro level Their strategies used in stoichiometric calculations were examined by an interview study of five Secondary Five students The connection between the submicro level and the symbolic level in learning stoichiometry would be also probed into Results indicate that the failure of stoichiometry learning may due to the disconnection in different levels of representation and students generated strategy one portion reasoning An implication for teaching and learning is that teachers should use diagrams at the submicro level in the teaching of stoichiometry Such diagrams should aim to help students building connections across the three levels of representation and enhancing students conceptual understanding in stoichiometry DOI 10 5353 th\_b5396387 Subjects Chemistry Study and teaching Secondary

**Chemical Calculations: an Educational Study of Stoichiometric Calculations in Secondary School Chemistry**  
Onno de Jong,1990

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