

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**

## *Gas Laws*

Chemists found that there were relationships among temperature, volume, pressure, and quantity of a gas that could be described mathematically. This chapter deals with Boyle's law, Charles's law, Gay-Lussac's law, the combined gas law, and Dalton's law of partial pressures. These laws have one condition in common. They all assume that the molar amount of gas does not change. In other words, these laws work correctly only when no additional gas enters a system and when no gas leaks out of it. Remember also that a law describes a fact of nature. Gases do not "obey" laws. The law does not dictate the behavior of the gas. Rather, each gas law describes a certain behavior of gas that occurs if conditions are right.

### **BOYLE'S LAW**

Robert Boyle, a British chemist who lived from 1627 to 1691 formulated the first gas law, now known as Boyle's law. This law describes the relationship between the pressure and volume of a sample of gas confined in a container. Boyle found that gases compress, much like a spring, when the pressure on the gas is increased. He also found that they "spring back" when the pressure is lowered. By "springing back" he meant that the volume increases when pressure is lowered. It's important to note that Boyle's law is true only if the temperature of the gas does not change and no additional gas is added to the container or leaks out of the container.

Boyle's law states that *the volume and pressure of a sample of gas are inversely proportional to each other at constant temperature*. This statement can be expressed as follows.

Proportionality symbol. It means "is proportional to."

Proportionality constant

$$V \propto \frac{1}{P} \quad \text{and} \quad PV = k \quad \text{or} \quad V = k \frac{1}{P}$$

According to Boyle's law, when the pressure on a gas is *increased*, the volume of the gas *decreases*. For example, if the pressure is doubled, the volume decreases by half. If the volume quadruples, the pressure decreases to one-fourth of its original value.

# Chemfile Mini Guide To Gas Laws

**WJ Hussar**



## **Chemfile Mini Guide To Gas Laws:**

## Unveiling the Magic of Words: A Report on "**Chemfile Mini Guide To Gas Laws**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Chemfile Mini Guide To Gas Laws**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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