



Chemical Changes

Read the passage and answer the following questions:



1. What are the products of this chemical reaction?

Mercury and oxygen

2. What is the reactant? Mercury oxide

3. What is the physical state of the reactant and the products?

The reactant is a solid, the products are a liquid and a gas.

4. According to the law of conservation of mass, if 100g of mercury oxide is heated, how much mercury and oxygen will be produced?

100g



5. What products are formed? Mercury and oxygen

6. What mass of reactant is used? 100g

7. What are the physical states of the reactant and the products?

Solid, liquid and gas

8. What is the physical state of the product? Liquid and gas

9. What is the mass of the products? 100g

Mercury and oxygen

10. How many moles of product are formed? 1 mole

Read the passage and answer the following questions:

11. How many moles of H_2O are produced if 1 mole of H_2 reacts with 1 mole of O_2 ?

2 moles of H_2O are produced

Chemical Changes Reinforcement Answers

Richard Bailey



Chemical Changes Reinforcement Answers:

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