

3. CHEMICAL KINETICS

Multiple Choice Type Questions

- 1 Unit of rate of reaction is—
(a) Mole per litre
(b) Mole per second
(c) Mole per litre per second
(d) Mole per litre 2 per second. Ans.(c)
- 2 Unit of rate constant for first order reaction is—
(a) $\text{mol L}^{-1} \text{s}^{-1}$,
(c) $\text{mol}^{-1} \text{L s}^{-1}$,
(b) s^{-1} ,
(d) None of these. Ans.(b)
- 3 Unit of rate constant for second order reaction, κ is—
(a) s^{-1} ,
(c) $\text{mol}^{-2} \text{L}^2 \text{s}^{-1}$,
(b) $\text{mol}^{-1} \text{L s}^{-1}$,
(d) $\text{mol L}^{-1} \text{s}^{-1}$. Ans.(b)
- 4 Unit of rate constant of third order reaction, κ is—
(a) s^{-1} ,
(c) $\text{mol}^{-2} \text{L}^2 \text{s}^{-1}$,
(b) $\text{mol}^{-1} \text{L s}^{-1}$,
(d) $\text{mol L}^{-1} \text{s}^{-1}$. Ans.(c)
- 5 Unit of rate constant of a zero order reaction is—
(a) s^{-1} ,
(c) $\text{L}^2 \text{mol}^{-2} \text{s}^{-1}$,
(b) $\text{mol L}^{-1} \text{s}^{-1}$,
(d) $\text{L mol}^{-1} \text{s}^{-1}$. Ans.(b)
- 6 The role of catalyst is to change—
(a) Gibbs energy,
(c) activation energy,
(b) enthalpy of reaction,
(d) equilibrium constant. Ans.(c)
- 7 Radioactive disintegration is an example of—
(a) Zero Order,
(c) Second Order,
(b) First order,
(d) Third Order. Ans.(b)
- 8 The unit of rate and rate constant are same for a reaction ?
(a) Zero Order,
(c) Second Order,
(b) First order,
(d) Third Order. Ans.(a)
- 9 The rate constant for a first order reaction is $2 \times 10^{-2} \text{min}^{-1}$ the half life period of reaction is—
(a) 69.3 min,
(c) 17.37 min,
(b) 34.65 min,
(d) 3.46 min. Ans.(b)
- 10 During the course of the chemical reaction the rate of reaction—
(a) Increases with time,
(b) Remains constant with time,
(c) Decreases with time,
(d) Irregular trend with time. Ans.(c)
- 11 The unit of rate constant for the reaction $2\text{H}_2 + 2\text{NO} \rightarrow 2\text{H}_2\text{O} + \text{N}_2$, which has rate $R = k [\text{H}_2] [\text{NO}]^2$ is—
(a) $\text{mol L}^{-1} \text{s}^{-1}$,
(c) $\text{mol}^{-2} \text{L}^2 \text{s}^{-1}$,
(b) s^{-1} ,
(d) mol L^{-1} . Ans.(c)

Chemical Kinetics Answers Multiple Choice Questions

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Ross, David B. Bailey**



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