

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
  - How many atoms of Fe are in one formula unit
  - How many sulfate ions can be attached to the iron atom
  - The charge on each Fe ion
  - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
  - Change the number of moles represented in the formula
  - Change the charges on the other ions in the compound
  - Change the formula so that it no longer represents the compound
  - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**  
**If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite**
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.  
**If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.**
- Fill in the blanks in the table below

| Compound Name               | Formula                      |
|-----------------------------|------------------------------|
| Aluminum sulfide            | $\text{Al}_2\text{S}_3$      |
| Aluminum sulfite            | $\text{Al}_2(\text{SO}_3)_3$ |
| Lead II chloride            | $\text{PbCl}_2$              |
| Ammonium phosphite          | $(\text{NH}_4)_3\text{PO}_3$ |
| Hydroiodic acid             | $\text{HI}$                  |
| Dinitrogen pentoxide        | $\text{N}_2\text{O}_5$       |
| Copper II carbonate         | $\text{CuCO}_3$              |
| Sodium sulfide              | $\text{Na}_2\text{S}$        |
| Ammonium phosphate          | $(\text{NH}_4)_3\text{PO}_4$ |
| Tin IV sulfite              | $\text{Sn}(\text{SO}_3)_2$   |
| Nitrous acid                | $\text{HNO}_2$               |
| Magnesium perchlorate       | $\text{Mg}(\text{ClO}_4)_2$  |
| Iron II nitrate             | $\text{Fe}(\text{NO}_3)_2$   |
| Iron III nitrate            | $\text{Fe}(\text{NO}_3)_3$   |
| Cobalt II oxide             | $\text{CoO}$                 |
| Sulfur trioxide             | $\text{SO}_3$                |
| Boron pentafluoride         | $\text{BF}_5$                |
| Tetraphosphorus octasulfide | $\text{P}_4\text{S}_8$       |
| Tetraiodide nonoxide        | $\text{I}_4\text{O}_9$       |

- Name the following compound and identify as molecular, ionic or acid
  - $\text{Ca}(\text{HCO}_3)_2$  **calcium bicarbonate - I**
  - $\text{Fe}(\text{ClO}_2)_3$  **Iron III chlorite - I**
  - $\text{Fe}(\text{ClO}_4)_2$  **Iron II chlorate - I**
  - $\text{N}_2\text{O}_5$  **dinitrogen pentoxide - M**

# Chemical Names And Formulas Review

**Benjamin Lillard**



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Publisher Description      **Practical Druggist and Pharmaceutical Review of Reviews** , 1918      **Chemistry** Hall, 1995-12

**Exercises in General Chemistry** Horace Grove Deming, Saul Bryan Arenson, 1924      **The American Perfumer and Essential Oil Review (1906)** , 1929      The American Perfumer and Essential Oil Review , 1922      *Modern Chemistry* , 1990      Basic Chemical Principles Edward I. Peters, 1988      *Prentice Hall Chemistry* , 2000      **Introduction to the Chemistry of Life** Peter P. Berlow, Donald J. Burton, Joseph Isaac Routh, 1982      **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      Essentials of Chemistry Dennis D. Staley, Antony C. Wilbraham, 1984

Chemist and Druggist , 1924      Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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