#### Six Types of Chemical Reaction Worksheet Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

a) 2 NaBr + 1 Ca(OH)₂ → 1 CaBr₂ + 2 NaOH

Type of reaction: double displacement

b) 2 NH<sub>3</sub>+ 1 H<sub>2</sub>SO<sub>4</sub> → 1 (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>

Type of reaction: synthesis

c) 4 C<sub>5</sub>H<sub>9</sub>O + 27 O<sub>2</sub> → 20 CO<sub>2</sub> + 18 H<sub>2</sub>O

Type of reaction: combustion

d) 3 Pb + 2 H<sub>3</sub>PO<sub>4</sub> → 3 H<sub>2</sub> + 1 Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

Type of reaction: single displacement

e) 1 Li<sub>3</sub>N + 3 NH<sub>4</sub>NO<sub>3</sub> → 3 LiNO<sub>3</sub> + 1 (NH<sub>4</sub>)<sub>3</sub>N

Type of reaction: double displacement

f) 3 HBr + 1 Al(OH)<sub>3</sub> → 3 H<sub>2</sub>O + 1 AlBr<sub>3</sub>

Type of reaction: acid-base (double displacement)

g) Na<sub>3</sub>PO<sub>4</sub> + 3 KOH → 3 NaOH + K<sub>3</sub>PO<sub>4</sub>

Reaction Type \_\_double displacement\_\_\_\_

Reaction Type \_\_double displacement\_\_\_\_

h) MgCl₂ + Li₂CO₃ → MgCO₃ + 2 LiCl

Reaction Type \_\_\_\_complete combustion\_\_\_

i) C<sub>6</sub>H<sub>12</sub> + 9 O<sub>2</sub> → 6 CO<sub>2</sub> + 6 H<sub>2</sub>O

Reaction Type \_\_\_\_single displacement \_\_\_\_

j) Pb + FeSO<sub>4</sub> → PbSO<sub>4</sub> + Fe

Reaction Type \_\_\_\_decomposition\_\_\_\_\_

k) CaCO<sub>2</sub> → CaO + CO<sub>2</sub>

Reaction Type \_\_\_synthesis\_\_\_\_\_

P<sub>4</sub> + 3 O<sub>2</sub> → 2 P<sub>2</sub>O<sub>3</sub>

Reaction Type \_\_\_\_double displacement\_\_\_\_

m) 2 RbNO<sub>3</sub> + BeF<sub>2</sub> → Be(NO<sub>3</sub>)<sub>2</sub> + 2 RbF

Reaction Type single displacement

n) 2 AgNO<sub>3</sub> + Cu → Cu(NO<sub>3</sub>)<sub>2</sub> + 2 Ag

Reaction Type \_\_\_\_complete combustion\_\_\_\_

O<sub>3</sub>H<sub>6</sub>O + 4 O<sub>2</sub> → 3 CO<sub>2</sub> + 3 H<sub>2</sub>O

Reaction Type \_\_\_synthesis\_\_\_\_

p) 2 C<sub>5</sub>H<sub>5</sub> + Fe → Fe(C<sub>5</sub>H<sub>5</sub>)<sub>2</sub>

Reaction Type single displacement

q) SeCl<sub>0</sub> + O<sub>2</sub> → SeO<sub>2</sub> + 3Cl<sub>2</sub>

Reaction Type \_\_\_\_double displacement\_\_\_\_

r) 2 Mgl<sub>2</sub> + Mn(SO<sub>3</sub>)<sub>2</sub> → 2 MgSO<sub>3</sub> + Mnl<sub>4</sub>

s) O<sub>3</sub> -> O + O<sub>2</sub>

Reaction Type \_\_\_\_decomposition\_\_\_\_\_

t) 2 NO<sub>2</sub> → 2 O<sub>2</sub> + N<sub>2</sub>

Reaction Type \_\_\_\_\_decomposition\_\_\_\_\_

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