

## Chemistry & Chemical Thermodynamics

### EXAM PRACTICE TES; Q&A

#### QUESTIONS 1

(1) An ideal gas,  $C_p = \left(\frac{7}{2}\right) R$  and  $C_v = \left(\frac{5}{2}\right) R$  undergoes a cycle consisting of the following mechanically reversible steps:

(a) An adiabatic compression from state A at  $P_1, V_1$  and  $T_1$  to state B at  $P_2, V_2, T_2$

(b) An isobaric expansion from state B at  $P_2, V_2, T_2$  to state C, at  $P_3 = P_2, V_3, T_3$

(c) An adiabatic expansion from state C at  $P_3, V_3, T_3$  to state D at  $P_4, V_4, T_4$

(d) A constant-volume expansion process from state D at  $P_4, V_4, T_4$  to state A at  $P_1, V_1, T_1$ .

(i) Sketch this cycle on a P V diagram

(ii) On the basis of one mole of ideal gas, what is the heat  $Q_{B \rightarrow C}$  absorbed in step (b), and the heat  $Q_{D \rightarrow A}$  rejected in step (d)?

(iii) If the thermal efficiency is defined as the ratio of the network output to the heat absorbed, show that thermal efficiency is

$$\eta = 1 + \frac{Q_{D \rightarrow A}}{Q_{B \rightarrow C}}$$

# Chemical Thermodynamics Practice Questions

**L Towner**



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