Name:

Date: -

STOICHIOMETRIC CALCULATION

1. How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction?

$$2 \operatorname{LiBr} (aq) + \operatorname{Cl}_2 (g) \rightarrow 2 \operatorname{LiCl} (aq) + \operatorname{Br}_2 (g)$$

2. Hydrogen peroxide breaks down, releasing oxygen, in the following reaction:

$$2 H_2O_2(aq) \rightarrow 2 H_2O(I) + O_2(g)$$

a. What mass of oxygen is produced when 1.840 mol of H₂O₂ decomposes?

b. What mass of water is produced when this reaction produces 5 mol of O₂?

Calculate the mass of silver bromide produced from 22.5 g of silver nitrate in the following reaction:

$$2 \text{ AgNO}_3 (aq) + \text{MgBr}_2 (aq) \rightarrow 2 \text{ AgBr} (s) + \text{Mg(NO}_3)_2 (aq)$$

4. Potassium chlorate is sometimes decomposed in the laboratory to generate oxygen. The reaction is:

$$2 \text{ KCIO}_3 (s) \rightarrow 2 \text{ KCI } (s) + 3 \text{ O}_2 (g)$$

What mass of KCIO₃ do you need to produce 0.5 mol of O₂?

<u>Chemical Calculations And Stoichiometry Study Island</u> <u>Answers</u>

Didier Musso

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