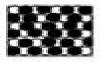
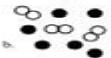
- This is practice Do NOT cheat yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- DO NOT USE A CALCULATOR. AP Chem does not allow the use of a calculator for the MC part of the exam, so it is time to start practicing without one. You may use ONLY a periodic table.
- Work only for about 1.5 minutes per question. Time yourself. It is important that you practice working for speed.
- The number of moles of NHs molecules in 34.0 g of NHs.
 - 20. 1
 - b. 2
 - e. 3
 - d. 5
- The number of moles in 3.01 × 10³⁴ atoms of carbon atoms
 - a. 2
 - bs. 3-
 - c. 5
 - d. 8
- The number of moles of S atoms in 6.02 = 10²³ S₀ molecules
 - in.
 - b. 2
 - G. 2
 - d. 8
- Which of the following contain approximately 3.01 × 10²³ particles
 - a. 32.0 g of oxygen molecules.
 - b. 16.0 g of oxygen molecules
 - c. 4.0 g of belium atoms
 - d. 23 g of sodium atoms.
- Which of the following contain(s) the same number of atoms as 48.6 g of magnesium
 - a. 56 g of nitrogen molecules
 - b. 48.6 g of calcium atoms
 - c. 4 moles of marnesium atoms
 - d. 1.2 × 10²⁴ atoms of carbon.
- All of the following quantities have a mass of 79.9 g EXCEPT
 - a. I mole of bromine atoms
 - 6.02 × 10²⁹ becoming atoms
 - c. 3.01 × 10²³ Br: molecules
 - d. 0.5 mole of bromine atoms
- Which of the following choices contains 0.5 mole of sodium ions?
 - n. 106 g sodium carbonate
 - b. 53 g of sodium carbonate
 - c. 26.5 g of sodium carbonate
 - d. 3.0 × 10²³ mole of sodium carbonate



 The diagram above represents solid carbon dioxide, better known as dry ice. Which diagram below best represents Some of the dry ice after it sublimes into a gas?









- If 1.5 × 10⁽¹⁾ atoms of an unknown element have a mass of 12 g, what is the molar mass of the element closest to?
 - a. 3.0 g mol 1
 - b. 12 g mol⁻¹
 - e. 24 g mol-1
 - d. 48 g mol-1
- 10. When the following chemical equation is bulanced using the lowest possible whole numbers, what is the sum of the stoichiometric coefficients?

- a. 66
- b. 81
- et 1600
- A 20
- 11. When the following chemical equation is bulanced using the lowest possible whole numbers, what is the sum of the stoichiometric coefficients?

 $H_3PO_4 + Mg(OH)_2 \rightarrow Mg_3(PO_4)_2 + H_2O$

- m. 4.
- b. 6
- c. 11
- di 12

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Thomas Evangelist

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